

# Sulfur clock lab

## Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

|   | Reactants | Reagents                                            | Apparatus             |
|---|-----------|-----------------------------------------------------|-----------------------|
| A | 1 mL      | 0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$ | Reaction setup        |
| B | 1.0 mL    | 0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$ | 0.1 M, 0.01 M, 0.01 M |
| C | 1.0 mL    | 0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$ | 0.1 M, 0.01 M, 0.01 M |
| D | 1.0 mL    | 0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$ | 0.1 M, 0.01 M, 0.01 M |

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the calculated amount of  $\text{H}_2\text{O}_2$  solution to the reaction mixture using a graduated cylinder.
6. Add exactly 1.0 mL of 0.1 M  $\text{KSCN}$  solution to the reaction mixture.
7. Prepare and use the standard solutions.
8. Prepare solutions A, B, C, D and E as shown:
  - a. Transfer 1.0 mL of 0.1 M  $\text{H}_2\text{O}_2$  solution to the reaction mixture.
  - b. Add 1.0 mL of 0.1 M  $\text{KSCN}$  solution to the reaction mixture.
  - c. Add 1.0 mL of 0.1 M  $\text{KSCN}$  solution to the reaction mixture.
  - d. Add 1.0 mL of 0.1 M  $\text{KSCN}$  solution to the reaction mixture.
9. Transfer solution A into the graduated cylinder as 1.0 mL of  $\text{H}_2\text{O}_2$ . Do not draw liquid out with any effort at all.

# Lab Sulfur Clock Answers

**Albert Emmett Lawrence**



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