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 $3\text{KClO}_3 \rightarrow 3\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{K} + 3\text{S} + 2\text{Al}_2\text{O}_3$
 $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$
 $8\text{C} + 5\text{O}_2 \rightarrow 2\text{C}_3\text{O}_2 + 4\text{CO}$

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web oct 2 2019 let's start by balancing the oxygen atoms to do this make the oxygen atoms as six on either side of the chemical equation $2\text{Fe}_2\text{O}_3 \rightarrow 4\text{Fe} + 3\text{O}_2$ on towards the next step now step 3 once you are done with this check out if the number of atoms for the other elements is also equal on both the sides

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 $\text{N}_2 + \text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr} + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr} + 3\text{SNO}_2 + 3\text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_x$
 $2\text{ZnCl}_x \rightarrow 2\text{Cu}$ how many moles of Ce ZnCl_2 ZnCl_x will be produced from 23.0 g of Ce 23.0 g of Zn

assuming $\text{CeCl}_2 \cdot 2\text{H}_2\text{O}$ is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{HBr}$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
br using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{KBr} + \text{CO}_2 + \text{H}_2\text{O}$
 $\text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{O}_3 + \text{SO}_2 + \text{NaBr}$
 SnF_2

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 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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same thing goes with the O_3 i am so lost

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