- (a) Anion is formed by the gain of electrons. Thus the numbers of electrons are more than protons. The effective positive charge in the nucleus is less, so less inward pull is experienced. Hence the size expands. So the size of an atom is greater than the size of parent atom.
- (b) Since Argon has stable octet configuration, so due to the inter- electronic repulsions the effect of nuclear pull over the valence shell electrons cannot be seen which results in the bigger size.
- (c) Since size of Bromine is bigger than chlorine, so it becomes more difficult for Br atoms to attract electrons. Thus, Cl is more reactive than Br.

Question 17:

Which element has:

- (a) two shells, both of which are completely filled with electrons?
- (b) the electronic configuration 2, 8, 3?
- (c) a total of three shells with five electrons in its valence shell?
- (d) a total of four shells with two electrons in its valence shell?
- (e) twice as many electrons in its second shell as in its first shell?

Solution 17:

- (a) Neon
- (b) Aluminum
- (c) Phosphorus
- (d) Calcium
- (e) Carbon

Question 18:

Name

- (a) An alkali metal in period 3 and halogen in period 2.
- (b) The noble gas with 3 shells.
- (c) The non-metals present in period 2 and metals in period 3.
- (d) The element of period 3 with valency 4
- (e) The element in period 3 which does not form oxide
- (f) The element of lower nuclear charge out of Be and Mg.
- (g) Which has higher E.A. fluorine or Neon.
- (h) Which has maximum metallic character Na, Li or K.

Solution 18.

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Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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Table of Contents Mastering Chemistry Chapter 7 Answers

- 1. Understanding the eBook Mastering Chemistry Chapter 7 Answers
 - The Rise of Digital Reading Mastering Chemistry Chapter 7 Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mastering Chemistry Chapter 7 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Chapter 7 Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mastering Chemistry Chapter 7 Answers
 - Personalized Recommendations
 - Mastering Chemistry Chapter 7 Answers User Reviews and Ratings
 - Mastering Chemistry Chapter 7 Answers and Bestseller Lists
- 5. Accessing Mastering Chemistry Chapter 7 Answers Free and Paid eBooks
 - Mastering Chemistry Chapter 7 Answers Public Domain eBooks
 - Mastering Chemistry Chapter 7 Answers eBook Subscription Services
 - Mastering Chemistry Chapter 7 Answers Budget-Friendly Options
- 6. Navigating Mastering Chemistry Chapter 7 Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Mastering Chemistry Chapter 7 Answers Compatibility with Devices
 - Mastering Chemistry Chapter 7 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Chapter 7 Answers
 - Highlighting and Note-Taking Mastering Chemistry Chapter 7 Answers
 - Interactive Elements Mastering Chemistry Chapter 7 Answers
- 8. Staying Engaged with Mastering Chemistry Chapter 7 Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Mastering Chemistry Chapter 7 Answers
- 9. Balancing eBooks and Physical Books Mastering Chemistry Chapter 7 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Chapter 7 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Chapter 7 Answers
 - Setting Reading Goals Mastering Chemistry Chapter 7 Answers
 - \circ Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Chapter 7 Answers
 - Fact-Checking eBook Content of Mastering Chemistry Chapter 7 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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