

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Mastering Chemistry Answer Guide

Harry E. Peery



Mastering Chemistry Answer Guide:

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Central Science James C. Hill, 1991 **A Guide to the Government Certificate Examinations. With Questions and Specimen Answers** R. J. Griffiths (LL.D.), 1871 **Chemistry, Student Solutions Manual** James E. Brady, Joel W. Russell, John R. Holum, 2000-01-26 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet *Student's Guide to Brown and LeMay, Chemistry, the Central Science, 2nd Edition* James C. Hill, 1981 **A Laboratory Manual of General Chemistry** William Crowell Bray, Ludwig Rosenstein, 1916

Study Guide & Solutions Manual to Accompany Organic Chemistry Mary H. Bailey, 1998 **Student's Guide to Brown and LeMay Chemistry** James C. Hill, 1988 *A Laboratory Manual of General Chemistry for Use in Colleges* William Crowell Bray, Wendell Mitchell Latimer, 1921 *The Success Manual for General Chemistry* Elizabeth Kean, Catherine Middlecamp, 1986 **Chemistry** James C. Hill, 2003 This book assists students through the text material with chapter overviews learning objectives review of key terms cumulative chapter review quizzes and self tests Included are answers to all Student Guide exercises Chapter summaries are correlated to those in the Instructor's Resource Manual

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