

# Chapter 4: Supramolecular Chemistry – Questions

1. Assign the oxidation number to each atom in the following elements, molecules and ionic substances:

- a)  $\text{CO}_2$       b)  $\text{CO}$       c)  $\text{HNO}_3$       d)  $\text{Ca(NO}_3)_2$       e)  $\text{H}_2\text{SO}_4$       f)  $\text{Ag}$   
 $\text{C} = +4$        $\text{C} = +2$        $\text{N} = +5$        $\text{Ca} = +2$        $\text{S} = +6$        $\text{Ag} = 0$   
 $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{N} = +5$        $\text{O} = -2$        $\text{Ag} = 0$
- g)  $\text{Fe}$       h)  $\text{BaCl}_2$       i)  $\text{CCl}_4$       j)  $\text{HBr}$       k)  $\text{HNO}_2$       l)  $\text{O}_2$   
 $\text{Fe} = 0$        $\text{Ba} = +2$        $\text{C} = +4$        $\text{Br} = -1$        $\text{N} = +3$        $\text{O} = 0$   
 $\text{Cl} = -1$        $\text{O} = -2$        $\text{H} = +1$        $\text{O} = -2$        $\text{O} = -2$

2. Assign the oxidation number to each atom in the following ions:

- a)  $\text{NO}^+$       b)  $\text{CO}_3^{2-}$       c)  $\text{HNO}_2$       d)  $\text{NO}_2^+$       e)  $\text{NO}_2^-$       f)  $\text{NO}_3^-$   
 $\text{N} = +2$        $\text{C} = +3$        $\text{N} = +3$        $\text{N} = +2$        $\text{N} = +3$        $\text{N} = +5$   
 $\text{O} = +1$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$

3. Write the following reactions identifying if it is an oxidation-reduction reaction. If it is a redox reaction identify the oxidising agent, the reducing agent, the substance being oxidised and the substance being reduced. Identify the oxidation number change to support these observations:



Redox rxn.  $\text{Cu}$  oxidised (oxidising agent);  $\text{H}_2$  as  $\text{H}^+$   $\longrightarrow$   $\text{H}_2$   
 $\text{HCl}$  reduced (red. reducing agent);  $\text{H}_2$  as  $\text{H}^+$   $\longrightarrow$   $\text{H}_2$



Not a redox rxn.  $\text{Ca}$  as  $\text{Ca}^{2+}$  is both oxidised and reduced



Redox rxn.  $\text{H}_2$  oxidised (oxidising agent);  $\text{H}_2$  as  $\text{H}^+$   $\longrightarrow$   $\text{H}_2$   
 $\text{H}_2$  reduced (red. reducing agent);  $\text{H}_2$  as  $\text{H}^+$   $\longrightarrow$   $\text{H}_2$



Not a redox reaction. No change in ox. #



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**Yan Bai**



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