

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

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Table of Contents Mastering Chemistry Study Guide

1. Understanding the eBook Mastering Chemistry Study Guide
 - The Rise of Digital Reading Mastering Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Study Guide
 - Personalized Recommendations
 - Mastering Chemistry Study Guide User Reviews and Ratings
 - Mastering Chemistry Study Guide and Bestseller Lists
5. Accessing Mastering Chemistry Study Guide Free and Paid eBooks

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- Mastering Chemistry Study Guide Budget-Friendly Options
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 - Interactive Elements Mastering Chemistry Study Guide
- 8. Staying Engaged with Mastering Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Study Guide
- 9. Balancing eBooks and Physical Books Mastering Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Study Guide
 - Setting Reading Goals Mastering Chemistry Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Study Guide
 - Fact-Checking eBook Content of Mastering Chemistry Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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