

Chapter 2 Answers

Practice 2-1

1. 2 2. -15 3. -14 4. -17 5. -41 6. 5 7. 19.7
8. -16.2 9. -7.6 10. $-\frac{1}{2}$ 11. $\frac{1}{3}$ 12. $-\frac{5}{12}$ 13. $1\frac{2}{3}$
14. $-2\frac{1}{2}$ 15. $-2\frac{1}{3}$ 16. 1.9 17. -0.99 18. 1.2 19. 33
20. 7 21. -7 22. -0.9 23. -0.7 24. -5 25. 5
26. -18 27. 1 28. -6 29. $\frac{5}{12}$ 30. $-2\frac{1}{3}$
31. $\begin{bmatrix} 3 & 1 \\ 0 & 2 \end{bmatrix}$ 32. $\begin{bmatrix} 0.4 \\ -0.4 \\ 1.1 \end{bmatrix}$
33. -18°F 34. their own 11-yd line 35. \$170.53 36. -39 ft

Practice 2-2

1. 7 2. -16 3. -12 4. -8 5. 43 6. -49 7. -21.4
8. 14.6 9. -9 10. 26.4 11. 12 12. -10.6 13. $-\frac{1}{2}$
14. -1 15. $\frac{1}{2}$ 16. -18 17. 12 18. -5.9 19. 24 20. 10.5
21. -0.99 22. 3 23. 9 24. -3 25. 9 26. -3 27. 3
28. 17 29. -8 30. -19 31. -7 32. -7 33. -8
34. $\begin{bmatrix} -8 & 1 \\ 5 & -4 \end{bmatrix}$ 35. $\begin{bmatrix} -0.9 & -1.7 \\ -2.1 & -6.3 \end{bmatrix}$
36. 29°F 37. 29,310 ft
38. -\$205.72 39. their own 35-yd line

Practice 2-3

1. -16 2. 54 3. 81 4. -32 5. -48 6. 196 7. 48 8. 6
9. 4 10. 120 11. -49 12. -243 13. -4 14. -2 15. 15
16. -125 17. 4 18. 112 19. $-\frac{4}{3}$ 20. -32 21. 49
22. -200 23. -20 24. 256 25. -11 26. 32 27. 0
28. -4 29. $-\frac{7}{4}$ 30. 16 31. 2 32. 91 33. 64 34. -120
35. -7 36. 3 37. 64 38. -15 39. -5 40. -15 41. 4
42. 72 43. -27 44. -1019 45. -15 46. -4 47. 108
48. 256

Practice 2-4

1. $2x + 12$ 2. $-40 + 5b$ 3. $-4x + 28$ 4. $-15c + 21$
5. $-7.5a - 12.5$ 6. $-3d + 12$ 7. $-9 + 12d$ 8. $4b - \frac{7}{5}$
9. $19.2x - 12.6$ 10. $10.5x - 28$ 11. $4x + 28$
12. $-5a + 10$ 13. $8 - 10d$ 14. $-2d + 22$ 15. $-2b - 5$
16. $-8c + 32$ 17. $-4 + 2b$ 18. $6a - 18$ 19. $8x + 32$
20. $-5b + 25$ 21. $3f + 6$ 22. $11h - 25$ 23. $d - 21$
24. $1 + 8x$ 25. $2h + 4$ 26. $8 + 2y$ 27. $-a - 2$
28. $3w + 12$ 29. $1.2d - 2$ 30. $-2d + 6$ 31. $5x + 12$
32. $6a + 4$ 33. $3y - 15$ 34. $-b + 20$ 35. $2d + 6$
36. $0.8x + 1.6$ 37. $6b - 18$ 38. $6w - 4$ 39. $x - 2$
40. $2x + 7$ 41. $9 + 10c$ 42. $1 + \frac{2}{5}d$ 43. $15x + 60$

44. $2w + 2$ 45. $8x - 9$ 46. $2x - 15$ 47. $3x - 36$
48. $-18 - 6d$ 49. $5(x + 6)$ 50. $2(y - 8)$
51. $-15(x - 5)$ 52. $\frac{32}{y + 12}$ 53. $-8(4 - w)$
54. $(x + 9)(7 - x)$

Practice 2-5

1. Comm. Prop. of Add. 2. Comm. Prop. of Add.
3. Ident. Prop. of Mult. 4. Distributive Prop.
5. Assoc. Prop. of Mult. 6. Inverse Prop. of Mult.
7. Distributive Prop. 8. Comm. Prop. of Add.
9. Assoc. Prop. of Add. 10. Inverse Prop. of Add.
11. Comm. Prop. of Add. 12. Assoc. Prop. of Mult.
13. Ident. Prop. of Add. 14. Comm. Prop. of Add.
15. Distributive Prop. 16. Mult. Prop. of Zero
17. Assoc. Prop. of Add. 18. Comm. Prop. of Mult.
19. Comm. Prop. of Mult. 20. Comm. Prop. of Add.
21a. Distributive Prop. 21b. Comm. Prop. of Add.
21c. Assoc. Prop. of Add. 21d. Distributive Prop.
21e. addition 22a. Distributive Prop. 22b. def. of subtr.
22c. Comm. Prop. of Add. 22d. Distributive Prop.
22e. addition 22f. def. of subtr. 23a. Distributive Prop.
23b. Comm. Prop. of Add. 23c. Distributive Prop.
23d. addition 24. 80 25. 7200 26. 2400 27. 18
28. \$7 29. \$28 30. \$16

Practice 2-6

1. $\frac{4}{7}$ 2. $\frac{2}{7}$ 3. $\frac{5}{7}$ 4. $\frac{4}{25}$ 5. $\frac{3}{25}$ 6. $\frac{12}{25}$ 7. 0 8. $\frac{21}{25}$ 9. $\frac{8}{25}$
10. $\frac{99}{100}$ 11. $\frac{1}{100}$ 12. 23,760 13a. $\frac{2}{3}$ 13b. $\frac{3}{4}$
14. 60% or $\frac{3}{5}$ 15a. $\frac{5}{14}$ 15b. $\frac{5}{7}$ 15c. $\frac{3}{7}$ 16a. $\frac{4}{13}$
16b. $\frac{8}{13}$ 16c. $\frac{7}{13}$

Practice 2-7

- 1a. $\frac{7}{80}$ 1b. $\frac{7}{76}$ 1c. $\frac{1}{16}$ 1d. $\frac{1}{19}$ 1e. $\frac{1}{10}$ 1f. $\frac{2}{19}$ 1g. $\frac{49}{400}$
1h. $\frac{21}{190}$ 2. $\frac{7}{9}$ 3. $\frac{10}{13}$ 4. $\frac{4}{5}$ 5. $\frac{2}{3}$ 6a. $\frac{21}{256}$ 6b. $\frac{7}{80}$
6c. $\frac{49}{256}$ 6d. $\frac{7}{40}$ 6e. $\frac{1}{64}$ 6f. $\frac{1}{120}$ 6g. $\frac{7}{64}$ 6h. $\frac{7}{60}$
7. $\frac{2}{3}$ 8. $\frac{1}{14}$ 9. $\frac{2}{3}$ 10. $\frac{3}{5}$ 11. $\frac{4}{7}$ 12. $\frac{2}{3}$

Reteaching 2-1

1. -7 2. 17 3. 3 4. -10 5. -5 6. -3 7. 2 8. -1
9. -3.8 10. 7.6 11. -2.3 12. 21.2 13. 0.2 14. -10.3
15. -20 16. -6.3 17. -1 18. -9 19. 1 20. 9
21. 5.9 22. 0.9 23. -0.9 24. -5.9 25. 10.5 26. 3.7
27. -3.7 28. -10.5

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occurs at pH 7.0

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4.4 weak acid and strong base phenolphthalein changes color in the basic range 8.2-10.6

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partly because so many compounds can act as acids or bases

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a 25.0 mL sample of 0.100 M hydrochloric acid with 0.100 M sodium hydroxide the values of the pH measured after successive
additions of small amounts of NaOH are listed in the first column of this table and are graphed in

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an acidic solution of known concentration a standard solution is added to a basic alkaline solution of unknown concentration
until the solution is exactly neutralised i.e. there is only salt and water it is possible to calculate

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titration of an acid with a base or vice versa the indicator is a substance that can exist in two forms an

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web the above equation works only for neutralizations in which there is a 1:1 ratio between the acid and the base the
example below demonstrates the technique to solve a titration problem for a titration of sulfuric acid with sodium hydroxide

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web Oct 13 2022 an acid base titration is a quantitative analysis method used to determine the concentration of an acid or
base by neutralizing the acid or base with a known concentration standard solution the concentration of a solution can be
determined by knowing the acid and base dissociation constant

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web there are many types of titration when considering goals and procedures however the most common types of titration in

quantitative chemical analysis are redox titration and acid base titration titrations can be classified as acid base titrations
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web a titration experiment can be carried out to accurately measure the volume of substances that react in chemical reactions this technique is commonly used in neutralisation reactions and can also

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web there are two basic types of acid base titrations indicator and potentiometric in an indicator based titration you add another chemical that changes color at the ph equal to the equivalence point when the acid and base are in stoichiometric proportions

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web the titration of a weak acid with a strong base involves the direct transfer of protons from the weak acid to the hydroxide

ion the reaction of the weak acid acetic acid with a strong base naoh can be seen below in the reaction the

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web as seen in the chapter on the stoichiometry of chemical reactions titrations can be used to quantitatively analyze solutions for their acid or base concentrations in this section we will explore the underlying chemical equilibria that make acid base titrimetry a

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