

# Chemistry Unit 4 Notes

Please Read Main Notes First Before Reading These

## Topic II - Further Kinetics

- 1) Rate of a chemical equation: It is the change in concentration of reactants or products per unit time. The concentration of the reactants will always decrease over time and the concentration of the products will always increase over time.
- 2) Rate Law: The rate of a chemical reaction is proportional to the initial concentration of reactants raised to their appropriate powers.
- 3) Rate Determining Step: It is the slowest step in a multistep mechanism.
- 4) Order of reaction with respect to a particular reactant/ Partial Order: It is the number of molecules of a particular reactant which are involved upto and including the **rate determining step (RDS)** of a reaction mechanism.
- 5) Overall order of a reaction: It is the total number of reactant molecules which are involved upto and including the RDS of the mechanism, that is, the sum of the partial orders.
- 6) Rate Constant: In the rate equation, the proportionality constant, **k**, is called the rate constant.

$$Rate = k[A]^m[B]^n$$

- 7) Arrhenius Equation:

$$k = Ae^{-E_a/RT}$$

Where,

k = rate constant

A = frequency factor

$E_a$  = activation energy

R = molar gas constant [8.31 Jmol<sup>-1</sup>K<sup>-1</sup>]

T = temperature in Kelvin

- 8) Half-life: It is the time taken for a reactant to fall to half of its initial concentration.
- 9) The half-life for a 1st order reaction does not depend on the concentration of the reactant.

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