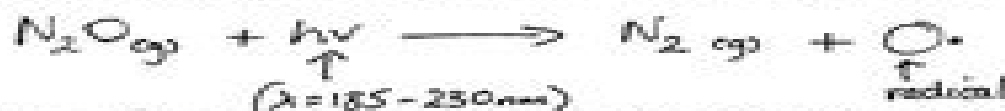


### F332 advance notice (June 2013) "predicted Q's"

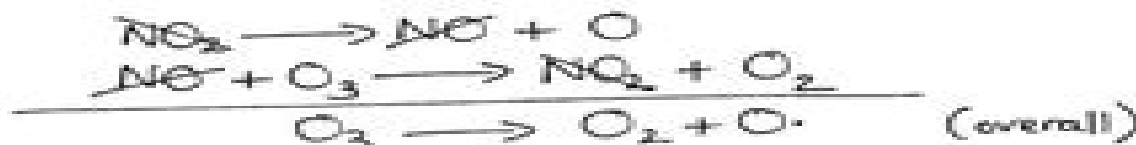
- How does photolysis play a part in the destruction of  $N_2O$  in the stratosphere?



[90% of  $N_2O$  loss in Stratosphere is this!]

\* THIS IS "PHOTODISSOCIATION".  $O\cdot$  helps to form tropospheric  $O_3$ .

- Use a series of equations to show how NO acts in the catalytic destruction of ozone?



This can be a bad thing at tropospheric level. NO can react with  $C_xH_y$ , preventing NO breaking down  $O_3$ . Build up of  $O_3$  in Trop. leads to photochemical smog!!

- What is meant by the term photolysis?

"photo" = LIGHT

"-lysis" = SPLITTING

Light energy drives chemical reactions

- Convert 100 parts per billion to a percentage, for the level of  $NO_x$  in urban areas

1 part per million  $\longrightarrow$   $\frac{1}{10,000}$  (divide by 10,000)

100 ppb = 0.1 ppm (1,000 ppb = 1 ppm)

$$\frac{0.1 \text{ ppm}}{10,000} = \underline{\underline{1 \times 10^{-5} \%}} \quad (\text{or } 0.00001 \%)$$

- Why is the  $N_2O_{(atm)} \rightleftharpoons N_2O_{(ocean)}$  not a true dynamic equilibrium?

ATMOSPHERE + OCEANS can't be regarded as ~~as~~ a "SEALED system".  
(gases "lost").

- Describe how photochemical smog can be harmful to:  
(a) the environment?

- crop damage

- compounds with  $C=C$  bonds are particularly vulnerable  
(rubbers, paints ...)

# June 2013 F332 Paper

**LP Steffe**



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