

**Mole Conversion Practice 1.0 – Answer Key****Objectives:** – Convert between moles, mass, and particles, using dimensional analysis.**Directions:** Solve the following mole conversion problems. Show all your work and include units!!!**1. Mass  $\leftrightarrow$  Moles**

- a. How many moles of sodium is 246.0 g of sodium? [10.70 mol]

$$246.0 \text{ g} = \left( \frac{1 \text{ mol}}{22.99 \text{ g}} \right) = \boxed{10.70 \text{ mol}}$$

- b. How many grams of silver are in 18.40 mol of silver? [1985 g]

$$18.40 \text{ mol} = \left( \frac{107.87 \text{ g}}{1 \text{ mol}} \right) = \boxed{1985 \text{ g}}$$

- c. How many grams of nitrogen dioxide are in 2.56 mol of nitrogen dioxide,
- $\text{NO}_2$
- ?

$$\text{Molar Mass of } \text{NO}_2 = 14.01 \text{ g} + (2)(16.00 \text{ g}) = 46.01 \text{ g/mol}$$

$$2.56 \text{ mol} = \left( \frac{46.01 \text{ g}}{1 \text{ mol}} \right) = \boxed{118 \text{ g}}$$

- d. How many moles of copper(II) phosphate are in 653.0 g of copper(II) phosphate?

$$\text{Molar Mass of } \text{Cu}_3(\text{PO}_4)_2 = (3)(63.55) + (2)(30.97) + (8)(16.00) = 380.39 \text{ g/mol}$$

$$653.0 \text{ g} = \left( \frac{1 \text{ mol}}{380.39 \text{ g}} \right) = \boxed{1.713 \text{ mol}}$$

**2. Moles  $\leftrightarrow$  Particles**

- a. Determine the number of atoms in 92.4 mol of aluminum. [
- $5.56 \times 10^{25}$
- atoms]

$$92.4 \text{ mol} = \left( \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \right) = \boxed{5.56 \times 10^{25} \text{ atoms}}$$

- b. How many moles are in
- $8.25 \times 10^{23}$
- molecules of iodine monochloride? [10.4 mol]

$$8.25 \times 10^{23} \text{ molecules} = \left( \frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ molecules}} \right) = \boxed{10.4 \text{ mol}}$$

# Mole Practice Problem Answer Key

**Joanne K. Price**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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