

Phet - Molecular Polarity

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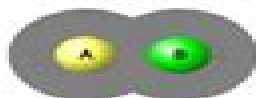
Part 1: TWO ATOMS

Select the "Two Atoms" tab at the top of the page and then make sure the following check boxes are turned on:

☒ Bond dipole, ☒ Partial Charges, and ☒ Electron Density.



1. In the "Electronegativity" boxes at the top, make Atom A "less" and Atom B "less".



The gray shading shows the distribution of electrons. When do you notice about the distribution of electrons when both atoms have LOW Electronegativity?

2. In the "Electronegativity" boxes at the top, keep Atom A "less" and gradually make Atom B "more".

a. What do you notice about the electron distribution as you make Atom B more electronegative?

b. Where are the partial charges located? Explain why.

c. Which way is the bond vector pointing?

3. In the "Electronegativity" boxes at the top, make Atom A "less" and gradually keep Atom B "more".

a. Describe the electron distribution and the partial charges when both atoms have high electronegativity.

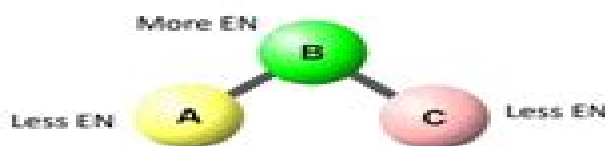
b. Explain why you think this happened.

Part 2: THREE ATOMS

Select the "Two Atoms" tab at the top of the page and then make sure the following check boxes are turned on:

☒ Bond dipole, ☒ Molecular Dipole, and ☒ Partial Charges.

Draw the **partial charges**, **bond dipoles**, and **molecular dipoles** for each situation.



Molecule Polarity Phet Lab

Sebastian Brünink



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