

**Unit 7 Chemical Reactions**  
**PRACTICE TEST/ REVIEW**

- \_\_\_\_\_ 1. A chemical reaction is a process in which
- a. products change into reactants
  - b. the law of conservation of mass applies
  - c. substances can change state
  - d. all of these
- \_\_\_\_\_ 2. During a chemical reaction,
- a. new elements are produced
  - b. atoms are rearranged
  - c. atoms are destroyed
  - d. elements are destroyed
- \_\_\_\_\_ 3. An equation is balanced by
- a. changing subscripts
  - b. adding coefficients
  - c. erasing elements as necessary
  - d. adding elements as necessary
- \_\_\_\_\_ 4. What are the reactants in the following chemical equation:
- $$\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$$
- a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 5. What are the products in the above equation?
- a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 6. The breaking down of a substance into two or simpler substances is \_\_\_\_\_
- a. Displacement
  - b. Synthesis
  - c. Decomposition
  - d. Double displacement
- \_\_\_\_\_ 7. Each substance to the right of the arrow in a chemical equation is a(n)
- a. Reactant
  - b. Catalyst
  - c. Product
  - d. Inhibitor
- \_\_\_\_\_ 8. According to the law of conservation of mass, if two atoms of hydrogen are used as a reactant, how many atoms of hydrogen must be part of the product?
- a. 4
  - b. 2
  - c. 1
  - d. 0
- \_\_\_\_\_ 9. What type of reaction is shown in the following chemical equation,  $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$ ?
- a. Displacement
  - b. Synthesis
  - c. Decomposition
  - d. Double displacement

# Modeling Chemistry Unit 7 Chemical Reactions Key

**RC Schank**

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## **Table of Contents Modeling Chemistry Unit 7 Chemical Reactions Key**

1. Understanding the eBook Modeling Chemistry Unit 7 Chemical Reactions Key
  - The Rise of Digital Reading Modeling Chemistry Unit 7 Chemical Reactions Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 7 Chemical Reactions Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 7 Chemical Reactions Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Chemical Reactions Key

- Personalized Recommendations
  - Modeling Chemistry Unit 7 Chemical Reactions Key User Reviews and Ratings
  - Modeling Chemistry Unit 7 Chemical Reactions Key and Bestseller Lists
5. Accessing Modeling Chemistry Unit 7 Chemical Reactions Key Free and Paid eBooks
    - Modeling Chemistry Unit 7 Chemical Reactions Key Public Domain eBooks
    - Modeling Chemistry Unit 7 Chemical Reactions Key eBook Subscription Services
    - Modeling Chemistry Unit 7 Chemical Reactions Key Budget-Friendly Options
  6. Navigating Modeling Chemistry Unit 7 Chemical Reactions Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Modeling Chemistry Unit 7 Chemical Reactions Key Compatibility with Devices
    - Modeling Chemistry Unit 7 Chemical Reactions Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Chemical Reactions Key
    - Highlighting and Note-Taking Modeling Chemistry Unit 7 Chemical Reactions Key
    - Interactive Elements Modeling Chemistry Unit 7 Chemical Reactions Key
  8. Staying Engaged with Modeling Chemistry Unit 7 Chemical Reactions Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modeling Chemistry Unit 7 Chemical Reactions Key
  9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Chemical Reactions Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Chemical Reactions Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Chemical Reactions Key
    - Setting Reading Goals Modeling Chemistry Unit 7 Chemical Reactions Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Chemical Reactions Key

- Fact-Checking eBook Content of Modeling Chemistry Unit 7 Chemical Reactions Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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