

NAME: Khang Tuong Lab Section: \_\_\_\_\_ Score: \_\_\_\_\_

## MOLECULAR MODELS Part I

In this experiment, you will use the Lewis structures that you draw to build models and explore the concepts of molecular shape, hybridization and bonding schemes and polarity. Show all  $[6N + 2] - [VE]$  work to aid in drawing Lewis Structures.

### Section I: $sp^3$ Hybridization:

Draw the Lewis structures on the table below. Include their VSEPR Structures and hybridization of the central atom. How many holes will the central atom require to accommodate the hybridized orbitals? 4 Build these four models and sketch a 3-D representation of them in the table.

Molecule	Lewis Structure	Model	VSEPR structure
CH <sub>4</sub> methane	$[6N + 2] - [VE] =$ $8 - 8 = 0$ $\rightarrow$ single bond only $\begin{array}{c} H \\   \\ H - C - H \\   \\ H \end{array}$		<p>Tetrahedral 109.5°</p>
CHCl <sub>3</sub> trichloromethane	$[6N + 2] - [VE] =$ $26 - 26 = 0$ $\rightarrow$ single bond only $\begin{array}{c} H \\   \\ :Cl: - C - :Cl: \\   \\ :Cl: \end{array}$		<p>Tetrahedral 109.5°</p>
CH <sub>2</sub> Cl <sub>2</sub> dichloromethane	$\begin{array}{c} :Cl: \\   \\ :Cl - C - H \\   \\ H \end{array}$		<p>Tetrahedral 109.5°</p>
CH <sub>2</sub> Br <sub>2</sub> dibromomethane	$\begin{array}{c} H \\   \\ :Br - C - Br: \\   \\ H \end{array}$		<p>Tetrahedral 109.5°</p>

# Modeling Chemistry U4 Lab V2

**Clifford Lane**



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