

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-E-A chart

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



1 mol  $\text{H}_2\text{O}$  : 18.02 g/mol

36 g  $\times \frac{1\text{ mol H}_2\text{O}}{18.02\text{ g}} = 1.99\text{ mol H}_2\text{O}$

1 mol  $\text{H}_2$  : 2.02 g

2 mol  $\text{H}_2 = \boxed{4.04\text{ g H}_2}$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol  $\text{NaCl} = 58.44\text{ g}$   
 $581\text{ g NaCl} \times \frac{1\text{ mol NaCl}}{58.44\text{ g}} = 9.94\text{ mol NaCl}$

1 mol  $\text{Na} = 22.99\text{ g}$

$9.94\text{ mol Na} \times 22.99\text{ g} = \boxed{228.52\text{ g Na}}$

# Modeling Chemistry Test U8 Testa V 2

**JS Bruner**



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## **Table of Contents Modeling Chemistry Test U8 Testa V 2**

1. Understanding the eBook Modeling Chemistry Test U8 Testa V 2
  - The Rise of Digital Reading Modeling Chemistry Test U8 Testa V 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Test U8 Testa V 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Test U8 Testa V 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Test U8 Testa V 2

- Personalized Recommendations
- Modeling Chemistry Test U8 Testa V 2 User Reviews and Ratings
- Modeling Chemistry Test U8 Testa V 2 and Bestseller Lists
- 5. Accessing Modeling Chemistry Test U8 Testa V 2 Free and Paid eBooks
  - Modeling Chemistry Test U8 Testa V 2 Public Domain eBooks
  - Modeling Chemistry Test U8 Testa V 2 eBook Subscription Services
  - Modeling Chemistry Test U8 Testa V 2 Budget-Friendly Options
- 6. Navigating Modeling Chemistry Test U8 Testa V 2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry Test U8 Testa V 2 Compatibility with Devices
  - Modeling Chemistry Test U8 Testa V 2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry Test U8 Testa V 2
  - Highlighting and Note-Taking Modeling Chemistry Test U8 Testa V 2
  - Interactive Elements Modeling Chemistry Test U8 Testa V 2
- 8. Staying Engaged with Modeling Chemistry Test U8 Testa V 2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry Test U8 Testa V 2
- 9. Balancing eBooks and Physical Books Modeling Chemistry Test U8 Testa V 2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry Test U8 Testa V 2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Test U8 Testa V 2
  - Setting Reading Goals Modeling Chemistry Test U8 Testa V 2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Test U8 Testa V 2

- Fact-Checking eBook Content of Modeling Chemistry Test U8 Testa V 2
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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