

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?



Before $\underline{8 \text{ moles}}$ $\xrightarrow{\text{excess}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-8 \text{ mol}}$ $\underline{-12 \text{ mol}}$ $\underline{+8 \text{ mol}}$ $\underline{+8 \text{ mol}}$

After — $\xrightarrow{\text{excess mol O}_2}$ $\underline{8 \text{ mol SO}_2}$ $\underline{8 \text{ mol H}_2\text{O}}$

2. Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?



Before $\underline{? \text{ mol}}$ $\xrightarrow{\text{excess}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-4 \text{ mol C}_3\text{H}_8}$ $\underline{-20 \text{ mol O}_2}$ $\underline{+12 \text{ mol CO}_2}$ $\underline{+16 \text{ mol H}_2\text{O}}$

After $\underline{0 \text{ mol C}_3\text{H}_8}$ $\xrightarrow{\text{excess mol O}_2}$ $\underline{12 \text{ mol CO}_2}$ $\underline{16 \text{ mol H}_2\text{O}}$

$$12 \text{ mol CO}_2 \times \frac{5 \text{ mol O}_2}{3 \text{ mol CO}_2}$$

3. Ammonia, NH_3 , for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?



Before $\xrightarrow{\text{excess mol}}$ $\underline{0.15 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\xrightarrow{\text{excess H}_2}$ $\underline{-0.15 \text{ mol}}$ $\underline{+0.3 \text{ mol}}$

After $\xrightarrow{\text{excess H}_2}$ $\underline{0 \text{ mol N}_2}$ $\underline{0.3 \text{ mol NH}_3}$

$$0.15 \text{ mol N}_2 \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} = 0.3 \text{ mol NH}_3$$

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ML Yell



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Table of Contents Modeling Chemistry Unit 8 Formula Practice

1. Understanding the eBook Modeling Chemistry Unit 8 Formula Practice
 - The Rise of Digital Reading Modeling Chemistry Unit 8 Formula Practice
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 8 Formula Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 8 Formula Practice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 8 Formula Practice

- Personalized Recommendations
 - Modeling Chemistry Unit 8 Formula Practice User Reviews and Ratings
 - Modeling Chemistry Unit 8 Formula Practice and Bestseller Lists
5. Accessing Modeling Chemistry Unit 8 Formula Practice Free and Paid eBooks
 - Modeling Chemistry Unit 8 Formula Practice Public Domain eBooks
 - Modeling Chemistry Unit 8 Formula Practice eBook Subscription Services
 - Modeling Chemistry Unit 8 Formula Practice Budget-Friendly Options
 6. Navigating Modeling Chemistry Unit 8 Formula Practice eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 8 Formula Practice Compatibility with Devices
 - Modeling Chemistry Unit 8 Formula Practice Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 8 Formula Practice
 - Highlighting and Note-Taking Modeling Chemistry Unit 8 Formula Practice
 - Interactive Elements Modeling Chemistry Unit 8 Formula Practice
 8. Staying Engaged with Modeling Chemistry Unit 8 Formula Practice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 8 Formula Practice
 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 8 Formula Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 8 Formula Practice
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Modeling Chemistry Unit 8 Formula Practice
 - Setting Reading Goals Modeling Chemistry Unit 8 Formula Practice
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Modeling Chemistry Unit 8 Formula Practice

- Fact-Checking eBook Content of Modeling Chemistry Unit 8 Formula Practice
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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