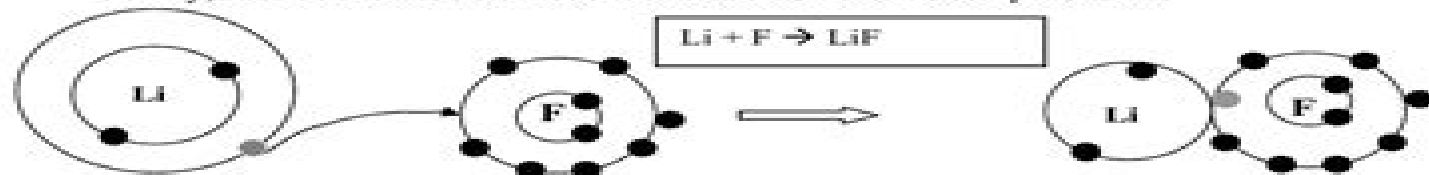


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For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms Valence electrons Electron transfer from/to each atom Ions formed in the product

Li
F

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$				
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$				
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$				
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$				
5) $\text{K} + \text{F} \Rightarrow \text{KF}$				

Ionic Bonding Worksheet

John S. Phillips, Cheryl Wistrom

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

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