

All answers in cells and lines are valued at 1 pt.  
each unless marked otherwise.

Name \_\_\_\_\_ Period \_\_\_\_\_

# Atoms and Bonds Chapter Test

## Section 1 Fill in the Table

Compound/Molecule	Metal + Nonmetal or Nonmetal + Nonmetal?	Ionic or Covalent Bond?	e <sup>-</sup> shared or transferred?
KI			
BN <sub>2</sub>			
CO <sub>2</sub>			
PbF <sub>2</sub>			

## Section 2 Periodic Symbol Analysis

28
<b>Ni</b>
58.693

Element name \_\_\_\_\_  
 Name English or Latin in origin? \_\_\_\_\_  
 Atomic number \_\_\_\_\_ Number of protons \_\_\_\_\_  
 Atomic mass \_\_\_\_\_ Number of electrons \_\_\_\_\_  
 Number of neutrons (Mass - # protons) \_\_\_\_\_  
 Metal, nonmetal, or noble gas? \_\_\_\_\_

## Section 3 Properties of Bonds Write "I" if the property describes an ionic bond, "C" if the property describes a covalent bond, and "B" if the property describes both kinds of bonds.

\_\_\_\_\_ High melting points \_\_\_\_\_ Weaker bond \_\_\_\_\_ Electrons shared  
 \_\_\_\_\_ Conducts electricity in water \_\_\_\_\_ Helps most atoms achieve the octet of a noble gas  
 \_\_\_\_\_ Polar or nonpolar molecules \_\_\_\_\_ Attraction of opposite charges  
 \_\_\_\_\_ Poor conductors of electricity \_\_\_\_\_ Crystal shape \_\_\_\_\_ Low melting points  
 \_\_\_\_\_ Attraction between a cation and an anion

## Section 4 Formula Writing Write the formula for each of the compounds below. Be sure to balance the charges with subscripts. (Hint: Parenthesized Roman numerals are the charges of transition metals)

- Silver sulfide \_\_\_\_\_
- Nickel (II) cyanide \_\_\_\_\_
- Aluminum phosphate \_\_\_\_\_
- Aluminum bromide \_\_\_\_\_
- Copper (I) sulfate \_\_\_\_\_
- Silver bromide \_\_\_\_\_

Cyanide = CN<sup>-</sup> Sulfide = S<sup>2-</sup> Phosphate = PO<sub>4</sub><sup>3-</sup> Silver = Ag<sup>+</sup> Sulfate = SO<sub>4</sub><sup>2-</sup> Aluminum = Al<sup>3+</sup>

# Ionic Bonds Chapter Test Answer Key

**John T. Moore, Richard H. Langley**



## **Ionic Bonds Chapter Test Answer Key:**

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