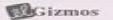
Activity A: Small atoms		Get the Gizmo ready:		
		<ul> <li>On the PERIODIC TABLE tab, select H (hydrogen).</li> <li>Select the ELECTRON CONFIGURATION tab.</li> <li>Click Reset.</li> </ul>	1 H 2	
Off	ganization are sho	rons are arranged in <b>orbitals</b> , <b>subshells</b> , and <b>shells</b> . These leaven by the boxes of the Gizmo. Each box represents an orbital d with letters ( $s$ , $p$ , $d$ , and $t$ ) and the shells are labeled with nur	. The	
Q	uestion: How are	electrons arranged in elements with atomic numbers 1 th	rough 10?	
1.	Infer: Based on	its atomic number, how many electrons does a hydrogen atom	have?	
2.	Arrange: The <b>Aufbau principle</b> states that electrons occupy the lowest-energy orbital. Click once in the <b>1s</b> box to add an electron to the only orbital in the <i>s</i> subshell of the first shell.			
	Click Check. What is the electron configuration of hydrogen?			
3.	Arrange: Click Next element to select helium. Add another electron to the 1s orbital. The arrows represent the spin of the electron. What do you notice about the arrows?			
	The Pauli exclusion principle states that electrons sharing an orbital have opposite spins.			
4.	Check your work	Check your work: Click Check. What is the electron configuration of helium?		
5.	Arrange: Click Next element and create electron configurations for lithium, beryllium, and boron. Click Check to check your work, and then list each configuration below:			
	Lithium:	Beryllium: Boron:		
6.	Arrange: Click Next element to select carbon. Add a second electron to the first 2p orbital.			
	Click Check. What feedback is given?			
7.	Rearrange: Hund's rule states that electrons will occupy an empty orbital when it is available in that subshell. Rearrange the electrons within the 2p subshell and click Check.			
	Is the configurat	ion correct now? 1s		
	Show the correc	t configuration in the boxes at right: 2s 2p		

(Activity A continued on next page)



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**Matthew Johll** 

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