

Ionic Compounds and Metals

Section 7.1 Ion Formation

pages 206–209

Section 7.1 Assessment

page 209

1. **Compare** the stability of a lithium atom with that of its ion, Li^+ .

The Li^+ ion is more stable because it has a complete octet.

2. **Describe** two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.

3. **Apply** Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.

4. **Summarize** ionic bond formation by correctly pairing these terms: *cation*, *anion*, *electron gain*, *electron loss*.

(anion, electron gain), (cation, electron loss)

5. **Apply** Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.

- a. nitrogen

$[\text{He}]2s^2 2p^3$; gain 3 electrons (3^- ion) or lose 5 electrons (5^+ ion)

- b. sulfur

$[\text{Ne}]3s^2 3p^4$; gain 2 electrons (2^- ion)

- c. barium

$[\text{Xe}]6s^2$; lose 2 electrons (2^+ ion)

- d. lithium

$[\text{He}]2s^1$; lose 1 electron (1^+ ion)

6. **Model** Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming Ca^{2+} , and that bromine gains one electron, forming Br^- . The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

Section 7.2 Ionic Bonds and Ionic Compounds

pages 210–217

Practice Problems

page 212

Explain how an ionic compound forms from these elements.

7. sodium and nitrogen

Three Na atoms each lose $1 e^-$, forming 1^+ ions. One N atom gains $3 e^-$, forming a 3^- ion. The ions attract, forming Na_3N .

$$3 \text{ Na ions } \left(\frac{1+}{\text{Na ion}} \right) + 1 \text{ N ion } \left(\frac{3-}{\text{N ion}} \right) \\ = 3(1+) + 1(3-) = 0$$

The overall charge on one formula unit of Na_3N is zero.

8. lithium and oxygen

Two Li atoms each lose $1 e^-$, forming 1^+ ions. One O atom gains $2 e^-$, forming a 2^- ion. The ions attract, forming Li_2O .

$$2 \text{ Li ions } \left(\frac{1+}{\text{Li ion}} \right) + 1 \text{ O ion } \left(\frac{2-}{\text{O ion}} \right) \\ = 2(1+) + 1(2-) = 0$$

The overall charge on one formula unit of Li_2O is zero.

Ionic Compounds And Metals Solutions Manual

John McMurry, Susan McMurry



Ionic Compounds And Metals Solutions Manual:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Table of Contents Ionic Compounds And Metals Solutions Manual

1. Understanding the eBook Ionic Compounds And Metals Solutions Manual
 - The Rise of Digital Reading Ionic Compounds And Metals Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Compounds And Metals Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Compounds And Metals Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Compounds And Metals Solutions Manual
 - Personalized Recommendations
 - Ionic Compounds And Metals Solutions Manual User Reviews and Ratings
 - Ionic Compounds And Metals Solutions Manual and Bestseller Lists

5. Accessing Ionic Compounds And Metals Solutions Manual Free and Paid eBooks
 - Ionic Compounds And Metals Solutions Manual Public Domain eBooks
 - Ionic Compounds And Metals Solutions Manual eBook Subscription Services
 - Ionic Compounds And Metals Solutions Manual Budget-Friendly Options
6. Navigating Ionic Compounds And Metals Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Compounds And Metals Solutions Manual Compatibility with Devices
 - Ionic Compounds And Metals Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Compounds And Metals Solutions Manual
 - Highlighting and Note-Taking Ionic Compounds And Metals Solutions Manual
 - Interactive Elements Ionic Compounds And Metals Solutions Manual
8. Staying Engaged with Ionic Compounds And Metals Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Compounds And Metals Solutions Manual
9. Balancing eBooks and Physical Books Ionic Compounds And Metals Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Compounds And Metals Solutions Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Compounds And Metals Solutions Manual
 - Setting Reading Goals Ionic Compounds And Metals Solutions Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Compounds And Metals Solutions Manual
 - Fact-Checking eBook Content of Ionic Compounds And Metals Solutions Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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