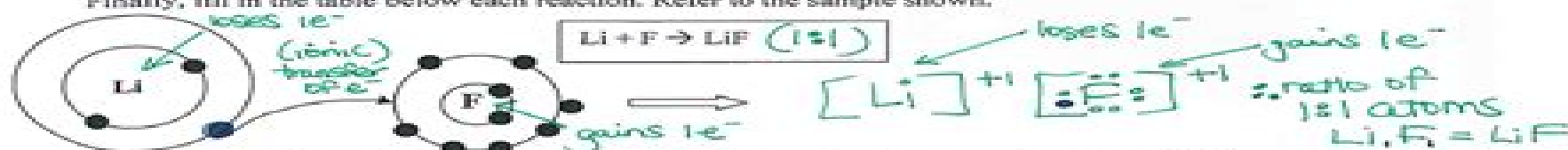


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Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$ $Li + \cdot\ddot{Cl}: \rightarrow [Li]^+ [\ddot{Cl}:]^{-1}$	Li	1 ve	donate $1e^-$	Li^+
	Cl	1 ve	accepts $1e^-$	Cl^{-1}
2) $Ca + O \Rightarrow CaO$ $Ca + \cdot\ddot{O}: \rightarrow [Ca]^{+2} [\ddot{O}:]^{-2}$	Ca	2 ve	donates $2e^-$	Ca^{+2}
	O	6 ve	accepts $2e^-$	O^{-2}
3) $Be + F \Rightarrow BeF_2$ $Be + \cdot\ddot{F}: \rightarrow [Be]^{+2} [\ddot{F}:]^{-1}$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
	F	7 ve	each F accepts $1e^-$ each	F^{-1}
4) $Mg + S \Rightarrow MgS$ $Mg + \cdot\ddot{S}: \rightarrow [Mg]^{+2} [\ddot{S}:]^{-2}$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
	S	6 ve	accepts $2e^-$	S^{-2}
5) $K + F \Rightarrow KF$ $K + \cdot\ddot{F}: \rightarrow [K]^+ [\ddot{F}:]^{-1}$	K	1 ve	donates $1e^-$	K^{+1}
	F	7 ve	accepts $1e^-$	F^{-1}

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