

Binary Ionic Compounds

Name: _____ Per: _____

Directions: First quickly scan the worksheet and circle any metals (as a symbol or as a name) that is a transition metal. Then either give the same or chemical formula as necessary for each problem below.

Remember that transition metals can have multiple oxidation states, so you are required to indicate the appropriate oxidation state in parenthesis when naming them. You must use the subscript and oxidation state of the anion to help determine the oxidation state of the transition metal.

- | | |
|--|-----------------------------------|
| 1. LiF _____ | 2. lithium chloride _____ |
| 3. Li ₂ O _____ | 4. lithium nitride _____ |
| 5. Li ₃ P _____ | 6. beryllium fluoride _____ |
| 7. BeO _____ | 8. beryllium sulfide _____ |
| 9. BF ₃ _____ | 10. boron chloride _____ |
| 11. BBr ₃ _____ | 12. boron oxide _____ |
| 13. BN _____ | 14. sodium fluoride _____ |
| 15. CuF _____ | 16. copper (II) chloride _____ |
| 17. Cu ₂ O _____ | 18. copper (II) oxide _____ |
| 19. Cu ₃ N _____ | 20. copper (II) nitride _____ |
| 21. Na ₂ O _____ | 22. sodium nitride _____ |
| 23. PbF ₂ _____ | 24. lead (IV) fluoride _____ |
| 25. PbS _____ | 26. lead (IV) sulfide _____ |
| 27. Pb ₃ N ₄ _____ | 28. lead (IV) oxide _____ |
| 29. FeF ₃ _____ | 30. iron (II) bromide _____ |
| 31. Fe ₂ O ₃ _____ | 32. chromium (VI) phosphide _____ |
| 33. FeP _____ | 34. aluminum fluoride _____ |
| 35. AlI ₃ _____ | 36. aluminum oxide _____ |
| 37. Mn ₂ O ₇ _____ | 38. manganese (IV) bromide _____ |
| 39. SnSe _____ | 40. tin (IV) oxide _____ |

Ionic Compound Worksheets With Answers

LL Leslie

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Ionic Compound Worksheets With Answers

1. Understanding the eBook Ionic Compound Worksheets With Answers
 - The Rise of Digital Reading Ionic Compound Worksheets With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Compound Worksheets With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Compound Worksheets With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Compound Worksheets With Answers
 - Personalized Recommendations

- Ionic Compound Worksheets With Answers User Reviews and Ratings
- Ionic Compound Worksheets With Answers and Bestseller Lists
- 5. Accessing Ionic Compound Worksheets With Answers Free and Paid eBooks
 - Ionic Compound Worksheets With Answers Public Domain eBooks
 - Ionic Compound Worksheets With Answers eBook Subscription Services
 - Ionic Compound Worksheets With Answers Budget-Friendly Options
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 - Ionic Compound Worksheets With Answers Compatibility with Devices
 - Ionic Compound Worksheets With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Compound Worksheets With Answers
 - Highlighting and Note-Taking Ionic Compound Worksheets With Answers
 - Interactive Elements Ionic Compound Worksheets With Answers
- 8. Staying Engaged with Ionic Compound Worksheets With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Compound Worksheets With Answers
- 9. Balancing eBooks and Physical Books Ionic Compound Worksheets With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Compound Worksheets With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Compound Worksheets With Answers
 - Setting Reading Goals Ionic Compound Worksheets With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Compound Worksheets With Answers
 - Fact-Checking eBook Content of Ionic Compound Worksheets With Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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