INDORE PUBLIC SCHOOL MID TERM EXAM PRACTICE PAPER

Class-XI

SUBJECT - CHEMISTRY (THEORY)

M.M.: 70

Time: 3 Hrs

General instructions: -

- (i) This question paper contains 26 questions. Write answers neatly and legibly. All questions are compulsory. This question paper has five sections: Section A, Section B, Section C, Section D and Section E.
- (ii) Question no. 1 to 5 are very short answer questions and carry 1 mark each.
- (iii) Question no. 6 to 10 are short answer questions and carry 2 marks each.
- (iv) Question no. 11 to 22 are also short answer questions and carry 3 marks each.
- (v) Question no. 23 is value based question and carry 4 marks.
- (vi) Question no. 24 to 26 are long answer questions and carry 5 marks each.
- (vii) Use log tables, if necessary. Use of calculators is not allowed.

SECTION-A [5 Questions (1 Mark Each), Total Marks: 5]

- Q.1 How many numbers of atoms are present in 1.4 gram of nitrogen gas?
- Q.2 Write the electronic configuration of Chromium and Copper? (At. No. are 24, 29 respectively)
- Q.3 What is the n_{factor} of KMnO₄ in acidic medium?
- Q.4 Define ionization energy?
- Q.5 Give hybridisation of XeO₂F₂ and CO₃²⁻.

SECTION-B [5 Questions (2 Marks Each), Total Marks: 10]

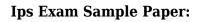
- Q.6 Nitrogen and Hydrogen gets combined to give Ammonia 0.5 mole of N₂ and 2.0 moles of H₂ are taken initially. Find:
 - i) What is limiting reagent?
 - ii) Weight of NH3 gets formed?
- Q.7 Define Aufbau rule and Pauli's Exclusion Principle.
- Q.8 What are the reasons behind diagonal relationship and write the elements which shows the diagonal relationship?
- Q.9 Among F and Cl, which one have the higher electron affinity. Give reason for your answer.
- Q.10 Which is more basic, NH₃ of PH₃? Give reason for your answer.

SECTION-C [12 Questions (3 Marks Each), Total Marks: 36]

- Q.11 a) Define Heisenberg Uncertainity Principle?
 - b) Minimum number of electron required to form a bond are?
- Q.12 Give the hybridisation and structure of PCl₃F₂.
- Q.13 The magnetic moment of an element is 3.8 B.M. (Z = 27).
 - a) No. of unpaired electrons are.
 - Write electronic configuration of the element.
- Q.14 A compound contains C = 71.23%, H = 12.95% and O = 15.81%. What is the empirical formula of the compound?
- Q.15 Define Molarity (M), Molality (m) and Normality (N).

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