

5 Descriptions of the bonding in three substances are given.

- substance 1 strong covalent bonds between atoms, permanent dipole-dipole attractions between molecules
- substance 2 strong covalent bonds between atoms, weak forces between molecules
- substance 3 strong covalent bonds between atoms, hydrogen bonding between molecules

Which compounds could be substances 1, 2 and 3?

	substance 1	substance 2	substance 3
<b>A</b>	CH <sub>3</sub> OH	Al <sub>2</sub> Cl <sub>6</sub>	CH <sub>2</sub> Cl <sub>2</sub>
<b>B</b>	Al <sub>2</sub> Cl <sub>6</sub>	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>4</sub>
<b>C</b>	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>4</sub>	CH <sub>3</sub> OH
<b>D</b>	CH <sub>4</sub>	CH <sub>3</sub> OH	H <sub>2</sub> O

6 Which type of bonding is **never** found in elements?

- A** covalent
- B** ionic
- C** metallic
- D** van der Waals' forces

7 Using the information in the table, what is the enthalpy change,  $q$ , for the reaction described?



standard enthalpy change	value / kJ mol <sup>-1</sup>
$\Delta H_{\text{f}}^\circ \text{ CsBr}(\text{s})$	+25.9
$\Delta H_{\text{f}}^\circ \text{ Cs}^+(\text{g})$	-276
$\Delta H_{\text{f}}^\circ \text{ Br}^-(\text{g})$	-335

- A** -636.9      **B** -585.1      **C** +585.1      **D** +636.9

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**EW Minium**



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## **Table of Contents May 2013 9701 A Level Chemistry Paper**

1. Understanding the eBook May 2013 9701 A Level Chemistry Paper
  - The Rise of Digital Reading May 2013 9701 A Level Chemistry Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying May 2013 9701 A Level Chemistry Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an May 2013 9701 A Level Chemistry Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from May 2013 9701 A Level Chemistry Paper
  - Personalized Recommendations
  - May 2013 9701 A Level Chemistry Paper User Reviews and Ratings
  - May 2013 9701 A Level Chemistry Paper and Bestseller Lists
5. Accessing May 2013 9701 A Level Chemistry Paper Free and Paid eBooks
  - May 2013 9701 A Level Chemistry Paper Public Domain eBooks
  - May 2013 9701 A Level Chemistry Paper eBook Subscription Services
  - May 2013 9701 A Level Chemistry Paper Budget-Friendly Options

6. Navigating May 2013 9701 A Level Chemistry Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - May 2013 9701 A Level Chemistry Paper Compatibility with Devices
  - May 2013 9701 A Level Chemistry Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of May 2013 9701 A Level Chemistry Paper
  - Highlighting and Note-Taking May 2013 9701 A Level Chemistry Paper
  - Interactive Elements May 2013 9701 A Level Chemistry Paper
8. Staying Engaged with May 2013 9701 A Level Chemistry Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers May 2013 9701 A Level Chemistry Paper
9. Balancing eBooks and Physical Books May 2013 9701 A Level Chemistry Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection May 2013 9701 A Level Chemistry Paper
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine May 2013 9701 A Level Chemistry Paper
  - Setting Reading Goals May 2013 9701 A Level Chemistry Paper
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of May 2013 9701 A Level Chemistry Paper
  - Fact-Checking eBook Content of May 2013 9701 A Level Chemistry Paper
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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