

F325 - June 2010, Q6e

17

- (e) Compound A is an orange ionic compound of chromium with the percentage composition by mass N, 11.11%; H, 3.17%; Cr, 41.27%; O, 44.45%. Compound A does not have water of crystallisation.

On gentle heating, compound A decomposes to form three products, B, C and water.

B is a green oxide of chromium with a molar mass of 152.0 g mol^{-1} .

C is a gas. At RTP, each cubic decimetre of C has a mass of 1.17 g.

In the steps below, show all your working.

- Calculate the empirical formula of compound A.
- Deduce the ions that make up the ionic compound A.
- Identify substances B and C.
- Write an equation for the decomposition of compound A by heat.

$$\begin{array}{r} \text{N} \\ 11.11\text{g} \\ \hline 14 \end{array}$$

$$\begin{array}{r} \text{H} \\ 3.17\text{g} \\ \hline 1 \end{array}$$

$$\begin{array}{r} \text{Cr} \\ 41.27\text{g} \\ \hline 52 \end{array}$$

$$\begin{array}{r} \text{O} \\ 44.45\text{g} \\ \hline 16 \end{array}$$

June 13 Chemistry F325

Wolfgang Guggemos

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

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Table of Contents June 13 Chemistry F325

1. Understanding the eBook June 13 Chemistry F325
 - The Rise of Digital Reading June 13 Chemistry F325
 - Advantages of eBooks Over Traditional Books
2. Identifying June 13 Chemistry F325
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 13 Chemistry F325
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 13 Chemistry F325
 - Personalized Recommendations

- June 13 Chemistry F325 User Reviews and Ratings
- June 13 Chemistry F325 and Bestseller Lists
- 5. Accessing June 13 Chemistry F325 Free and Paid eBooks
 - June 13 Chemistry F325 Public Domain eBooks
 - June 13 Chemistry F325 eBook Subscription Services
 - June 13 Chemistry F325 Budget-Friendly Options
- 6. Navigating June 13 Chemistry F325 eBook Formats
 - ePub, PDF, MOBI, and More
 - June 13 Chemistry F325 Compatibility with Devices
 - June 13 Chemistry F325 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 13 Chemistry F325
 - Highlighting and Note-Taking June 13 Chemistry F325
 - Interactive Elements June 13 Chemistry F325
- 8. Staying Engaged with June 13 Chemistry F325
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 13 Chemistry F325
- 9. Balancing eBooks and Physical Books June 13 Chemistry F325
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 13 Chemistry F325
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 13 Chemistry F325
 - Setting Reading Goals June 13 Chemistry F325
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 13 Chemistry F325
 - Fact-Checking eBook Content of June 13 Chemistry F325

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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