

EQUILIBRIUM TEST

MULTIPLE CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. It is so very large, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen trioxide, N_2O_3 , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



In a given closed system, the equilibrium constant for this reaction is 0.50. If 0.50 M of the reactant N_2O_3 is placed in a container at 0.500 M at this temperature, the product concentrations of NO_2 at equilibrium is

- A. 0.500 M
- B. 1.00 M
- C. 1.50 M
- D. 2.00 M

10) Chloroacetic acid, ClCH_2COOH , is known to react, H_2O , to form, methanoic acid, H_2COOH . H_2COOH is known to react with H_2O to form hydroxyacetic acid, HOCH_2COOH , and hydronium ions, H_3O^+ according to the equilibrium below:



A reaction that involves 2.0 g of chloroacetic acid is placed in a closed and held overnight. At that time the hydroxyacetic acid is found to be at equilibrium. In the morning, more change is made in the system. Which one of the following changes is most likely to result in an increase in the concentration of H_3O^+ ions?

- A. Change the volume.
- B. Bubble in more ClCH_2COOH gas.
- C. Add a few drops of perchloric acid solution.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Multiple Choice Concept Questions On Equilibrium

Greg Jacobs



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