

CUET Chemistry MCQs

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- (D) None of these
15. The number of tetrahedral and octahedral voids in hexagonal prismatic unit cell HCP is
(A) 8,4 (B) 2,1
(C) 12,6 (D) 6,12
16. Which of the following void is smallest in close packing?
(A) Tetrahedral (B) Octahedral
(C) Cubic (D) Square
17. The number of octahedral void in BCC structure is
(A) 0 (B) 1
(C) 2 (D) 4
18. Which of the following interstitial site is formed when the three closed packed spheres of one layer is put over three closed packed spheres of the second layer, their positions being inverted with respect to each other?
(A) Tetrahedral (B) Octahedral
(C) Rhombohedral (D) Tetragonal
19. In which crystal, alternate tetrahedral voids are occupied?
(A) NaCl (B) CaF₂
(C) ZnS (D) Na₂O
20. A metallic element crystallizes into a lattice containing a sequence of layers of ABABAB. Any packing of spheres leaves out voids in the lattice. Approximately what percentage by volume of this lattice is empty space?
(A) 74% (B) 26%
(C) 48% (D) 32%
21. Packing efficiency of an identical solid sphere is 74% in
(A) Simple cubic structure
(B) Monoclinic structure
(C) HCP structure
(D) BCC structure

22. Which of the following expression represents total volume of atoms present in a FCC unit cell of a metal? (r is the atomic radius)
(A) $\frac{20}{3}\pi r^3$ (B) $\frac{24}{3}\pi r^3$
(C) $\frac{12}{3}\pi r^3$ (D) $\frac{16}{3}\pi r^3$
23. In a BCC packing, the nearest neighbours lie along the
(A) edges of the cube
(B) face diagonal
(C) line joining the two opposite corners of the face.
(D) cube diagonal
24. The ionic radii of Rb⁺ and I⁻ are 1.46 and 2.16 Å, respectively. The most probable type of structure exhibited by it is
(A) CsCl type (B) NaCl type
(C) ZnS type (D) CaF₂ type
25. The radius of Ag⁺ ion is 126 pm while of I⁻ ion is 216 pm. The coordination number of Ag in AgI is
(A) 2 (B) 4
(C) 6 (D) 8
26. Structure shown here represents:

(A) Schottky defect
(B) Frenkel defect
(C) Metal excess defect
(D) None of these
27. Select the incorrect statement:
(A) Stoichiometry of crystal remains unaffected due to Schottky defect.

Multiple Choice Questions Of Solid State Chemistry

S Ashworth



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