

# How to Identify the Limiting & Excess Reactants/Reagents

Molar Mass:

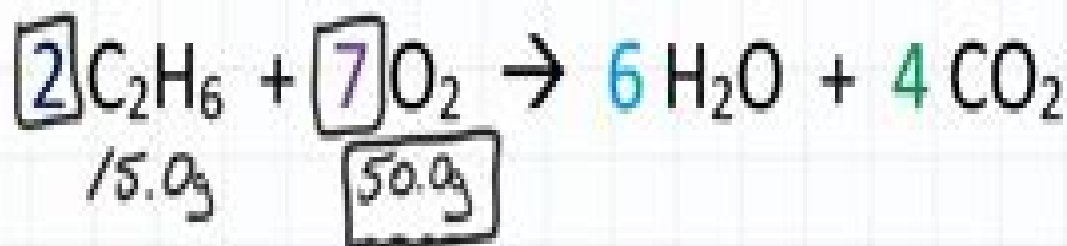
$\text{C}_2\text{H}_6 = 30.08 \text{ g/mol}$

$\text{O}_2 = 32.00 \text{ g/mol}$

$\text{H}_2\text{O} = 18.02 \text{ g/mol}$

1.) Calculate how much of one reactant will react with the other reactant

Ex. What is the limiting reactant when 15.0 g of  $\text{C}_2\text{H}_6$  react with 50.0 g of  $\text{O}_2$ ?



$$\overset{\text{A}}{\text{15.0g } \cancel{\text{C}_2\text{H}_6}} \left( \frac{1 \cancel{\text{mol C}_2\text{H}_6}}{30.08 \cancel{\text{g C}_2\text{H}_6}} \right) \left( \frac{7 \cancel{\text{mol O}_2}}{2 \cancel{\text{mol C}_2\text{H}_6}} \right) \left( \frac{32.00 \cancel{\text{g O}_2}}{1 \cancel{\text{mol O}_2}} \right) = \underline{\underline{55.9 \text{ g O}_2}} \quad \overset{\text{B}}{\text{O}_2 = \text{L.R.}}$$

$$\overset{\text{A}}{50.0 \cancel{\text{g O}_2}} \left( \frac{1 \cancel{\text{mol O}_2}}{32.00 \cancel{\text{g O}_2}} \right) \left( \frac{2 \cancel{\text{mol C}_2\text{H}_6}}{7 \cancel{\text{mol O}_2}} \right) \left( \frac{30.08 \cancel{\text{g C}_2\text{H}_6}}{1 \cancel{\text{mol C}_2\text{H}_6}} \right) = 13.4 \text{ g C}_2\text{H}_6 \quad \overset{\text{B}}{\text{C}_2\text{H}_6 = \text{E.R.}}$$

If reacting all of Reactant A requires **more** of Reactant B than is available, then **B** is the Limiting Reactant.

If reacting all of Reactant A requires **less** of Reactant B than is available, then **A** is the Limiting Reactant.

# Limiting Reactant Homework Answers On Chemfiesta

**K Payea**



### **Limiting Reactant Homework Answers On Chemfiesta:**

Limiting Reactant and Percent Yield, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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