

## Section A

Answer all questions in the spaces provided

0 1

Anhydrous magnesium chloride,  $\text{MgCl}_2$ , can absorb water to form the hydrated salt  $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}$



*← creating crystals directly is not really possible.*

0 1 - 1

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Virtually impossible to prevent dissolving.

0 1 - 2

Some enthalpies of solution are shown in Table 1.

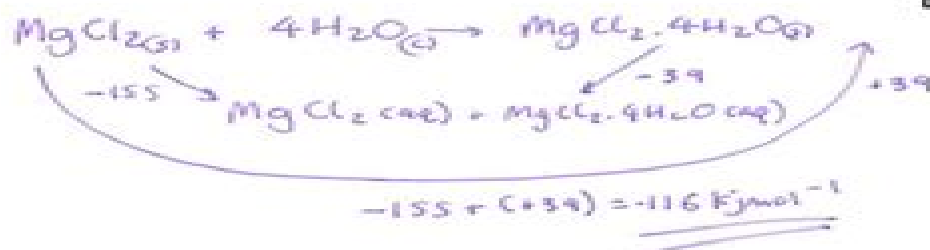
Table 1

Salt	Enthalpy of solution / $\text{kJ mol}^{-1}$
$\text{MgCl}_2(\text{s})$	-155
$\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$	-39

*making the salts aqueous!*

Calculate the enthalpy change for the absorption of water by  $\text{MgCl}_2(\text{s})$  to form  $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$ .

[2 marks]



Enthalpy change \_\_\_\_\_  $\text{kJ mol}^{-1}$



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**R Barnett**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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