

THE NATURE OF GASES

This section describes how the kinetic theory applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance.

Match the correct state of matter with each description of water by writing a letter on each line.

_____ 1. tap water

a. solid

_____ 2. Steam

b. liquid

_____ 3. Ice

c. gas

4. A substance may change from one state to another after a change in _____.

5. The energy an object has because of its motion is called _____.

6. Circle the letter of each sentence that is true about the assumption of the kinetic theory concerning gases.

- a. A gas is composed of particles with insignificant volume that are relatively far apart from each other.
- b. Strong attractive forces exist between particles of a gas.
- c. Gases tend to collect near the bottom of a container.
- d. The paths of uninterrupted travel of gas particles are relatively short because the particles are constantly colliding with each other or with other objects.

7. Is the following statement true or false? According to the kinetic theory, collisions between gas particles are perfectly elastic because kinetic energy is transferred without loss from one particle to another, and the total kinetic energy remains constant. _____

Nature Of Gases Section Review Answers

A Gutmann



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