

13.1

THE NATURE OF GASES

Section Review

Objectives

1. describe the characteristics of ideal gases and real gases
2. distinguish between an ideal gas and a real gas
3. define the compressibility factor, Z , and explain its significance

Vocabulary

- | | | |
|--------------|---------------------------|-----------------------|
| 1. ideal gas | 4. pressure | 7. pressure |
| 2. real gas | 5. compressibility factor | 8. ideal gas equation |
| 3. gas law | 6. volume | |

Part A Completion

Use the vocabulary terms in parentheses to complete each sentence. Write your answers in the spaces provided.

1. A(n) _____ gas obeys the _____ equation of state, and its behavior is described by the _____ law.
2. A(n) _____ gas is a gas that does not obey the _____ law.
3. A(n) _____ gas is a gas that does not obey the _____ law.
4. A(n) _____ gas is a gas that does not obey the _____ law.
5. A(n) _____ gas is a gas that does not obey the _____ law.
6. A(n) _____ gas is a gas that does not obey the _____ law.
7. A(n) _____ gas is a gas that does not obey the _____ law.
8. A(n) _____ gas is a gas that does not obey the _____ law.
9. A(n) _____ gas is a gas that does not obey the _____ law.
10. A(n) _____ gas is a gas that does not obey the _____ law.

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Part B True-False

Write **T** for true or **F** for false in the space provided. (If you are unsure, write **U**.)

1. _____ All gases are ideal gases.
2. _____ All gases are real gases.

Nature Of Gases Section Review Key

JR Anderson



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