

CHEMISTRY AND PHYSICS PRACTICALS

CHEMISTRY

- You are provided with ammonium chloride (NH_4Cl), sodium hydroxide pellets (NaOH), measuring cylinder, thermometer, stirring rod, distilled water, spatula.
 - Pour 50cm^3 of distilled water into one test tube.
 - Measure the initial temperature in the test tube and record in the table of results.
 - Add half a spatula of ammonium chloride (NH_4Cl) into the test tube and stir gently.
 - Measure the final temperature of the solution in the test tube and record the result.
 - Remove the thermometer from the test tube and rinse it with water.
 - Repeat the steps (a) to (d) with sodium hydroxide pellets and record the results.

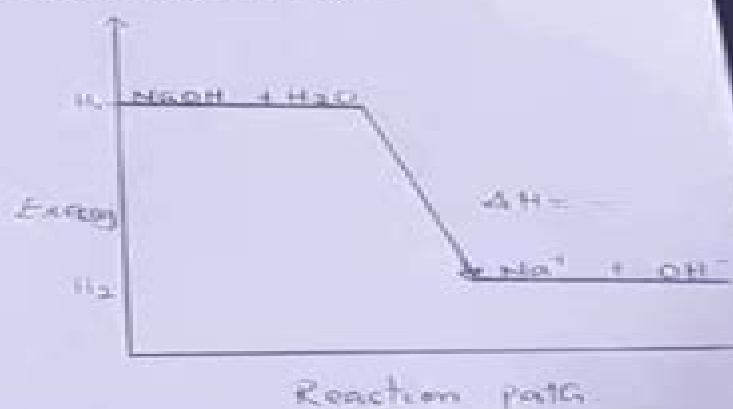
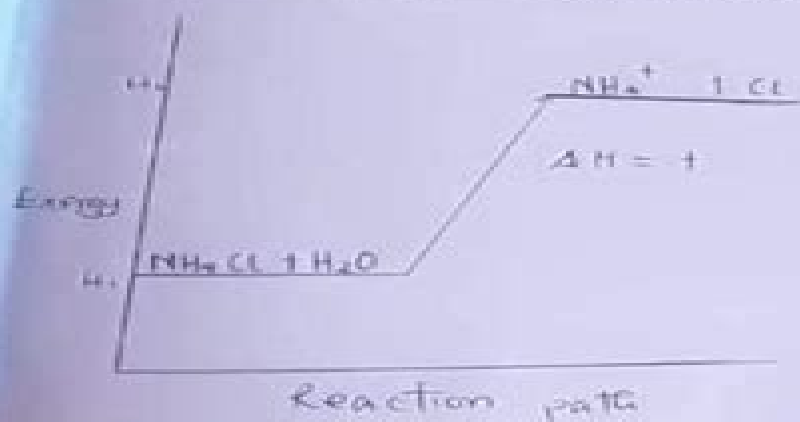
Solution	Initial ($^{\circ}\text{C}$) temperature	Final ($^{\circ}\text{C}$) temperature	Temperature change (final - initial)
NH_4Cl	28°C	25°C	-3°C
NaOH	28°C	45°C	17°C

- Classify the reactions as exothermic or endothermic.

Water + ammonium chloride Endothermic

Water + sodium hydroxide Exothermic

- Draw energy level diagrams to illustrate the above two reactions.



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Table of Contents Maneb Question Pracatical

1. Understanding the eBook Maneb Question Pracatical
 - The Rise of Digital Reading Maneb Question Pracatical
 - Advantages of eBooks Over Traditional Books
2. Identifying Maneb Question Pracatical
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Maneb Question Pracatical
 - User-Friendly Interface
4. Exploring eBook Recommendations from Maneb Question Pracatical
 - Personalized Recommendations
 - Maneb Question Pracatical User Reviews and Ratings
 - Maneb Question Pracatical and Bestseller Lists
5. Accessing Maneb Question Pracatical Free and Paid eBooks

- Maneb Question Pracatical Public Domain eBooks
- Maneb Question Pracatical eBook Subscription Services
- Maneb Question Pracatical Budget-Friendly Options
- 6. Navigating Maneb Question Pracatical eBook Formats
 - ePub, PDF, MOBI, and More
 - Maneb Question Pracatical Compatibility with Devices
 - Maneb Question Pracatical Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Maneb Question Pracatical
 - Highlighting and Note-Taking Maneb Question Pracatical
 - Interactive Elements Maneb Question Pracatical
- 8. Staying Engaged with Maneb Question Pracatical
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Maneb Question Pracatical
- 9. Balancing eBooks and Physical Books Maneb Question Pracatical
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Maneb Question Pracatical
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Maneb Question Pracatical
 - Setting Reading Goals Maneb Question Pracatical
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Maneb Question Pracatical
 - Fact-Checking eBook Content of Maneb Question Pracatical
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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