

# Stoichiometry - Mixed Conversions Answer Key

1. Use the equation below the answer the following questions:



- c. How many grams of  $\text{NH}_3$  are made from 31L of  $\text{N}_2$  gas?

$$\frac{31\text{L N}_2}{1} \times \frac{1\text{mol N}_2}{22.4\text{L N}_2} \times \frac{2\text{mol NH}_3}{1\text{mol N}_2} \times \frac{17\text{g mol NH}_3}{1\text{mol NH}_3} = 47.3\text{g NH}_3$$

- d. How many molecules of  $\text{H}_2$  are needed to produce 66g of  $\text{NH}_3$ ?

$$\frac{66\text{g NH}_3}{1} \times \frac{1\text{mol NH}_3}{17\text{g mol NH}_3} \times \frac{1\text{mol N}_2}{2\text{mol NH}_3} \times \frac{6.02 \times 10^{23}\text{ molecules N}_2}{1\text{mol N}_2} = 1.22 \times 10^{24}\text{ molecules N}_2$$

- e. How many moles of  $\text{N}_2$  are needed to completely react with 87L of  $\text{H}_2$  gas?

$$\frac{87\text{L H}_2}{1} \times \frac{1\text{mol H}_2}{22.4\text{L H}_2} \times \frac{1\text{mol N}_2}{3\text{mol H}_2} = 1.26\text{mol N}_2$$

2. Butane reacts with oxygen to form water and carbon dioxide.



- c. How many molecules of  $\text{H}_2\text{O}$  are produced with 20.1L of  $\text{O}_2$  gas that are completely used during the combustion of butane?

$$\frac{20.1\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{10\text{mol H}_2\text{O}}{13\text{mol O}_2} \times \frac{6.02 \times 10^{23}\text{ molecules H}_2\text{O}}{1\text{mol H}_2\text{O}} = 4.17 \times 10^{23}\text{ molecules H}_2\text{O}$$

- d. If 48g of  $\text{H}_2\text{O}$  are collected from the burning of butane, then how many moles of butane were reacted with oxygen?

$$\frac{48\text{g H}_2\text{O}}{1} \times \frac{1\text{mol H}_2\text{O}}{18\text{g H}_2\text{O}} \times \frac{2\text{mol C}_4\text{H}_{10}}{10\text{mol H}_2\text{O}} = 0.53\text{mol C}_4\text{H}_{10}$$

3. Potassium reacts with water to produce potassium hydroxide and hydrogen gas.



- c. How many grams of  $\text{H}_2$  will be produced if  $7.22 \times 10^{23}$  atoms of potassium react completely with water?

$$\frac{7.22 \times 10^{23}\text{ atoms K}}{1} \times \frac{1\text{mol K}}{6.02 \times 10^{23}\text{ atoms K}} \times \frac{1\text{mol H}_2}{2\text{mol K}} \times \frac{2\text{g H}_2}{1\text{mol H}_2} = 0.60\text{g H}_2$$

- d. How many atoms of K are needed to completely react with  $3.22 \times 10^{23}$  molecules of water?

$$\frac{3.22 \times 10^{23}\text{ molecules H}_2\text{O}}{1} \times \frac{1\text{mol H}_2\text{O}}{6.02 \times 10^{23}\text{ molecules H}_2\text{O}} \times \frac{2\text{mol K}}{1\text{mol H}_2\text{O}} \times \frac{6.02 \times 10^{23}\text{ atoms K}}{1\text{mol K}} = 3.22 \times 10^{23}\text{ atoms K}$$

# Mixed Stoichiometry Worksheet

**Clemens Wendtner**

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## **Table of Contents Mixed Stoichiometry Worksheet**

1. Understanding the eBook Mixed Stoichiometry Worksheet
  - The Rise of Digital Reading Mixed Stoichiometry Worksheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Stoichiometry Worksheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mixed Stoichiometry Worksheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Stoichiometry Worksheet
  - Personalized Recommendations

- Mixed Stoichiometry Worksheet User Reviews and Ratings
- Mixed Stoichiometry Worksheet and Bestseller Lists
- 5. Accessing Mixed Stoichiometry Worksheet Free and Paid eBooks
  - Mixed Stoichiometry Worksheet Public Domain eBooks
  - Mixed Stoichiometry Worksheet eBook Subscription Services
  - Mixed Stoichiometry Worksheet Budget-Friendly Options
- 6. Navigating Mixed Stoichiometry Worksheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Mixed Stoichiometry Worksheet Compatibility with Devices
  - Mixed Stoichiometry Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mixed Stoichiometry Worksheet
  - Highlighting and Note-Taking Mixed Stoichiometry Worksheet
  - Interactive Elements Mixed Stoichiometry Worksheet
- 8. Staying Engaged with Mixed Stoichiometry Worksheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mixed Stoichiometry Worksheet
- 9. Balancing eBooks and Physical Books Mixed Stoichiometry Worksheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mixed Stoichiometry Worksheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Stoichiometry Worksheet
  - Setting Reading Goals Mixed Stoichiometry Worksheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Stoichiometry Worksheet
  - Fact-Checking eBook Content of Mixed Stoichiometry Worksheet



- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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