Name: ______ Date: _____ Pd:____

MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.

$$_1$$
Al₂(SO₄)₃ + $_3$ Ca(OH)₂ \rightarrow $_2$ Al(OH)₃ + $_3$ CaSO₄
342.17 g/mol 74.10 g/mol 78.01 g/mol 136.15 g/mol

STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: W5.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

96 Al =
$$\frac{total\ mass\ Al}{molar\ mass\ Al_2(SO_4)_3}$$
 x 100 % = $\frac{53.96\ g}{342.17\ g}$ x 100 % = 15.77 %

STEP 4: Complete the following problems using a factor chart with units & chemicals. These problems could require ONE ratio. TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

How many grams of calcium sulfate are produced from reacting 2.59 moles of aluminum sulfate?

2.59 mol Al₂(SO₄)₃ x
$$\frac{3 mol CaSO_4}{1 mol Al_2(SO_4)_3}$$
 x $\frac{136 15 g CaSO_4}{1 mol CaSO_4}$ = 1057.89 g CaSO₄

moles A problem B grams B

Mixed Stoichiometry Answers

Richard Bailey

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