

2013

(May)

CHEMISTRY

(Major)

Course: 201

(Physical, Inorganic, Organic)

Full Marks: 80

Pass Marks: 32

Time: 3 hours

The figures in the margin indicate full marks for the questions

Write the answers to the separate Sections in separate books

SECTION – I

(Physical Chemistry)

(Marks: 26)

1. Choose the correct answer:

1×3=3

a) A thermodynamic state function is a quantity

- i. Used to determine heat changes.
- ii. Whose value is independent of path?
- iii. Used to determine pressure-volume work.
- iv. Whose value depends on temperature only.

b) The enthalpy of combustion of carbon is -394 kJ mol^{-1} . The heat evolved in the combustion of 6.02×10^{22} atoms of carbon is

- i. 3940 kJ
- ii. 394 kJ
- iii. 39.4 kJ
- iv. 0.394 kJ

c) The equilibrium constant for a reaction is 10. The value of standard free energy, ΔG° is

- i. -5.7 kJ mol^{-1}
- ii. 5.7 kJ mol^{-1}
- iii. -57 kJ mol^{-1}
- iv. 57 kJ mol^{-1}

UNIT – I

Answer any two of the following:

6×2=12

2. (a) Explain why we have to define the heat capacity of gases under constant pressure and constant volume condition.

2

(b) Thermodynamically show that for an ideal gas $C_p - C_v = R$.

3

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