Unit 5 - Chemical Reactions

- 13 Write complete balanced chemical equations for the following. Classify each reaction as combination, decomposition, single replacement, double replacement, neutralization, or combustion. Don't forget your subscripts!
- (a) potassium reacts with bromine

2 K(s) +1Brzu, -> 2 KBros)

(b) phosphoric acid is added to potassium hydrogen earbonate solution

11-3704001 + 3 KHCO36001 -> 342CO36001 + 1 K3PO4602)

Double replacement (Neutralization) (carbonate neutralization)

(c) magnesium is added to nickel (II) nitrate solution

Mages, + Ni(NO3)289) -> Mg(NO3)2000) + Nics)

Single Replacement.

(d) hydrochloric acid is added to ammonium hydroxide.

HCleg + NH40Heag -> NH4Cleg + H2Q1)

Double Replacement (neutralization)

(e) nickel is added to magnesium chloride solution

Nics + MgCl2(00) -> No Reaction

(f) copper (II) sulphate pentahydrate is produced in a combination reaction

CuSO461 + 5H2Ocg) -> CuSO4.5H2Ocs)

Combination

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Scott C. Dulebohn

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Treatise on Physical Chemistry Hugh Stott Taylor,1924 35th Aerospace Sciences Meeting & Exhibit ,1997

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