

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-E-A chart

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



1 mol  $\text{H}_2\text{O}$  : 18.02 g/mol

36 g  $\frac{1\text{ mol H}_2\text{O}}{18.02\text{ g}} = 1.99\text{ mol H}_2\text{O}$

1 mol  $\text{H}_2$  : 2.02 g

2 mol  $\text{H}_2 = \boxed{4.04\text{ g H}_2}$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol  $\text{NaCl} = 58.44\text{ g}$   
 $581\text{ g NaCl} \times \frac{1\text{ mol NaCl}}{58.44\text{ g}} = 9.94\text{ mol NaCl}$

1 mol  $\text{Na} = 22.99\text{ g}$

$9.94\text{ mol Na} \times 22.99\text{ g} = \boxed{228.52\text{ g Na}}$

# Modeling Chemistry U8 Ws 1 V2 0 Key

**J Dewey**



## **Modeling Chemistry U8 Ws 1 V2 0 Key:**

The book delves into Modeling Chemistry U8 Ws 1 V2 0 Key. Modeling Chemistry U8 Ws 1 V2 0 Key is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Modeling Chemistry U8 Ws 1 V2 0 Key, encompassing both the fundamentals and more intricate discussions.

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  - Chapter 1: Introduction to Modeling Chemistry U8 Ws 1 V2 0 Key
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  - Chapter 3: Modeling Chemistry U8 Ws 1 V2 0 Key in Everyday Life
  - Chapter 4: Modeling Chemistry U8 Ws 1 V2 0 Key in Specific Contexts
  - Chapter 5: Conclusion
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