Date _____ Pd ____

Chemistry – Unit 7 Worksheet 3

Write balanced chemical equations for the following reactions.

Note: All coefficients of 1 are omitted; in the first 10 items, the coefficients are all 1.

- Ammonia (NH₃) reacts with hydrogen chloride to form ammonium chloride.
 NH₂ + HCl → NH₄Cl
- Calcium carbonate decomposes upon heating to form calcium oxide and carbon dioxide.
 CaCO₂ → CaO + CO₂
- Barium oxide reacts with water to form barium hydroxide.
 BaO + H₂O → Ba(OH)₂
- Acetaldehyde (CH3CHO) decomposes to form methane (CH₄) and carbon monoxide.
 CH₂CHO → CH₄ + CO
- Zinc reacts with copper (II) nitrate to form zinc nitrate and copper.
 Zn + Cu(NO₃)₂ → Zn(NO₃)₂ + Cu
- Calcium sulfite decomposes when heated to form calcium oxide and sulfur dioxide.
 CaSO₃ → CaO + SO₂
- Iron reacts with sulfuric acid (H₂SO₄) to form iron (II) sulfate and hydrogen gas.
 Fe + H₂SO₄ → FeSO₄ + H₂
- A nitrogen containing carbon compound, C₂H₆N₂, decomposes to form ethane, C₂H₆, and nitrogen gas.
 C₂H₆N₂ → C₂H₆ + N₂
- Phosgene, COCl₂, is formed when carbon monoxide reacts with chlorine gas.
 CO + Cl₂ → COCl₂

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