

## Chemistry – Unit 7 Worksheet B

Write balanced chemical equations for the following reactions.

*Hydrogen gas (collected over water) is then vented to the fume hood. It is removed from the flask.*

1. Acetylene ( $\text{C}_2\text{H}_2$ ) reacts with water hydrogen chloride for three acetylene, a hydrogen gas.  

$$3\text{C}_2\text{H}_2 + 2\text{HCl} \rightarrow 3\text{C}_2\text{H}_4$$

2. Calcium carbonate is decomposed by heating. The three calcium oxide and carbon dioxide.  

$$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$$

3. Barium sulfate reacts with water for three barium hydroxide.  

$$\text{BaSO}_4 + 2\text{H}_2\text{O} \rightarrow \text{Ba(OH)}_2$$

4. Dichloromethane ( $\text{CH}_2\text{Cl}_2$ ) decomposes to form methane ( $\text{CH}_4$ ) and carbon monoxide.  

$$\text{CH}_2\text{Cl}_2 \rightarrow \text{CH}_4 + \text{CO}$$

5. Zinc reacting with copper (II) chloride for three zinc chloride and copper.  

$$\text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$$

6. Calcium sulfate decomposes when heated to form calcium oxide and sulfur dioxide.  

$$\text{CaSO}_4 \rightarrow \text{CaO} + \text{SO}_2$$

7. Iron reacting with sulfur gas ( $\text{Fe}_2\text{S}_3$ ) to form iron (II) sulfide and hydrogen gas.  

$$\text{Fe} + \text{H}_2\text{S}_3 \rightarrow \text{FeS}_2 + \text{H}_2$$

8. A nitrogen containing carbon compound,  $\text{C}_2\text{H}_5\text{N}_3$ , decomposes to form ethane,  $\text{C}_2\text{H}_6$ , and nitrogen gas.  

$$\text{C}_2\text{H}_5\text{N}_3 \rightarrow \text{C}_2\text{H}_6 + \text{N}_2$$

9. Phosphorus,  $\text{P}_4\text{O}_{10}$ , is formed when carbon monoxide reacts with phosphorus gas.  

$$\text{P}_4 + 10\text{CO} \rightarrow \text{P}_4\text{O}_{10} + 10\text{C}$$

# Modeling Chemistry Unit 7 Ws3 Answers

**Sabine Zange**



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## **Table of Contents Modeling Chemistry Unit 7 Ws3 Answers**

1. Understanding the eBook Modeling Chemistry Unit 7 Ws3 Answers
  - The Rise of Digital Reading Modeling Chemistry Unit 7 Ws3 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 7 Ws3 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 7 Ws3 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Ws3 Answers
  - Personalized Recommendations
  - Modeling Chemistry Unit 7 Ws3 Answers User Reviews and Ratings
  - Modeling Chemistry Unit 7 Ws3 Answers and Bestseller Lists
5. Accessing Modeling Chemistry Unit 7 Ws3 Answers Free and Paid eBooks
  - Modeling Chemistry Unit 7 Ws3 Answers Public Domain eBooks
  - Modeling Chemistry Unit 7 Ws3 Answers eBook Subscription Services
  - Modeling Chemistry Unit 7 Ws3 Answers Budget-Friendly Options
6. Navigating Modeling Chemistry Unit 7 Ws3 Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Modeling Chemistry Unit 7 Ws3 Answers Compatibility with Devices
  - Modeling Chemistry Unit 7 Ws3 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Ws3 Answers
    - Highlighting and Note-Taking Modeling Chemistry Unit 7 Ws3 Answers
    - Interactive Elements Modeling Chemistry Unit 7 Ws3 Answers
  8. Staying Engaged with Modeling Chemistry Unit 7 Ws3 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modeling Chemistry Unit 7 Ws3 Answers
  9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Ws3 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Ws3 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Ws3 Answers
    - Setting Reading Goals Modeling Chemistry Unit 7 Ws3 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Ws3 Answers
    - Fact-Checking eBook Content of Modeling Chemistry Unit 7 Ws3 Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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