

Chemistry Unit 8

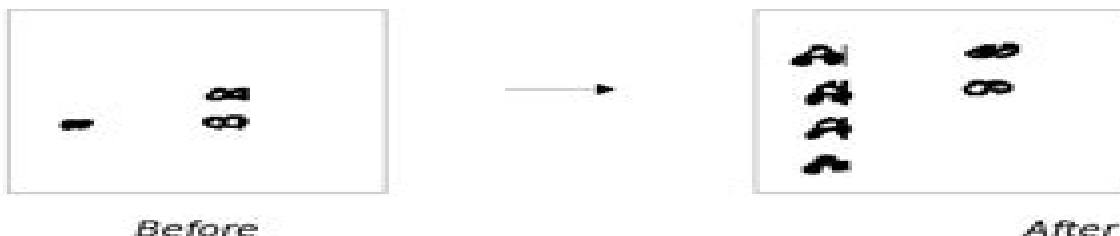
Worksheet 3: Adjusting to Reality - Limiting Reactant

1. Write the balanced equation for the reaction between hydrogen and oxygen.

Balanced Equation: $2\text{H}_2 + 1\text{O}_2 \rightarrow 2\text{H}_2\text{O}$

Suppose that 4 molecules of hydrogen gas and 4 molecules of oxygen gas react to form water.

Make a drawing that represents the reaction container before and after the reaction.



- 4 How many molecules of water can be produced?

Oxygen, Leftover molecules Which reactant is in excess? Why?

- 2 How many molecules of excess reactant are there?

Construct a Before-Change-After Table for this reactant mixture:

Bal. Equation: $2\text{H}_2 + 1\text{O}_2 \rightarrow 2\text{H}_2\text{O}$

| | | | |
|---------|----|----|----|
| Before: | 4 | 4 | 0 |
| Change: | -4 | -2 | +4 |
| After: | 0 | 2 | 4 |

According to the table you just made,

- 4 How many molecules of water can be produced?

Oxygen, 2 molecules left over post-reaction Which reactant is in excess? Why?

- 2 How many molecules of excess reactant are there?

Based on your two methods of analysis above, what determines how much product can be made from a particular reactant mix?

The determining factor is which of the reactants runs out first. This determines how much product can be made, even if another reactant is in excess.

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