

## **MIXED GAS LAWS WORKSHEET - SOLUTIONS**

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of  $O_2$  and 3.0 moles of  $N_2$  are placed in a 30.0 L tank at a temperature of  $25^\circ C$ , what will the pressure of the resulting mixture of gases be?  $25^\circ C = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is  $20.0^\circ C$ . By noon the temperature has risen to  $45.0^\circ C$ . What is the new volume of the balloon?

$$T_1 = 20.0^\circ C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ C = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in  $^\circ C$  of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ C$$

6)  $CaCO_3$  decomposes at  $1200^\circ C$  to form  $CO_2$  gas and  $CaO$ . If 25 L of  $CO_2$  are collected at  $1200^\circ C$ , what will the volume of this gas be after it cools to  $25^\circ C$ ?

$$T_1 = 1200^\circ C = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ C = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at  $20.0^\circ C$  is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is  $-20.0^\circ C$ ?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ C = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ C = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

# Mixed Gas Law Calculations Answers

**G Orfield**



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## **Table of Contents Mixed Gas Law Calculations Answers**

1. Understanding the eBook Mixed Gas Law Calculations Answers
  - The Rise of Digital Reading Mixed Gas Law Calculations Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Gas Law Calculations Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mixed Gas Law Calculations Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Gas Law Calculations Answers
  - Personalized Recommendations
  - Mixed Gas Law Calculations Answers User Reviews and Ratings
  - Mixed Gas Law Calculations Answers and Bestseller Lists
5. Accessing Mixed Gas Law Calculations Answers Free and Paid eBooks
  - Mixed Gas Law Calculations Answers Public Domain eBooks
  - Mixed Gas Law Calculations Answers eBook Subscription Services

- Mixed Gas Law Calculations Answers Budget-Friendly Options
- 6. Navigating Mixed Gas Law Calculations Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Mixed Gas Law Calculations Answers Compatibility with Devices
  - Mixed Gas Law Calculations Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mixed Gas Law Calculations Answers
  - Highlighting and Note-Taking Mixed Gas Law Calculations Answers
  - Interactive Elements Mixed Gas Law Calculations Answers
- 8. Staying Engaged with Mixed Gas Law Calculations Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mixed Gas Law Calculations Answers
- 9. Balancing eBooks and Physical Books Mixed Gas Law Calculations Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mixed Gas Law Calculations Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Gas Law Calculations Answers
  - Setting Reading Goals Mixed Gas Law Calculations Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Gas Law Calculations Answers
  - Fact-Checking eBook Content of Mixed Gas Law Calculations Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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