If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$V = \frac{nRT}{P} = \frac{7.7 \times 329 \times .0821}{.090tm} = 2310.9L$$

 If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

8) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have

$$N = \frac{RT}{PV}$$
 $\frac{.0821 \times 300}{.5a4m \times 25L} = 1.97mol$

If I have 21 moles of gas held at a pressure of 78 atm and a temperature of 900 K, what is the volume of the gas?

10) If I have 1.9 moles of gas held at a pressure of 5 atm and in a container with a volume of 50 liters, what is the temperature of the gas?

11) If I have 2.4 moles of gas held at a temperature of 97.0°C and in a container with a volume of 45 liters, what is the pressure of the gas?370

Mixed Gas Laws Answers

Christopher R. Hren, John T. Moore, Peter J. Mikulecky

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In some sort of inundated with monitors and the cacophony of immediate conversation, the profound power and psychological resonance of verbal artistry usually disappear in to obscurity, eclipsed by the constant assault of noise and distractions. Yet, situated within the lyrical pages of **Mixed Gas Laws Answers**, a charming perform of fictional brilliance that impulses with fresh emotions, lies an remarkable trip waiting to be embarked upon. Composed with a virtuoso wordsmith, that enchanting opus courses readers on a mental odyssey, softly revealing the latent possible and profound influence embedded within the elaborate internet of language. Within the heart-wrenching expanse of this evocative examination, we can embark upon an introspective exploration of the book is main themes, dissect their charming publishing design, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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