

# STOICHIOMETRY: MASS-MASS PROBLEMS

Name \_\_\_\_\_



How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose?

$$25\text{g KClO}_3 \left( \frac{1\text{ mol}}{122.5\text{g}} \right) \left( \frac{2\text{ mol KCl}}{2\text{ mol KClO}_3} \right) \left( \frac{74.5\text{g KCl}}{1\text{ mol KCl}} \right) = 13.2\text{g KCl}$$



How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

$$50\text{g N}_2 \left( \frac{1\text{ mol N}_2}{28\text{g N}_2} \right) \left( \frac{3\text{ mol H}_2}{1\text{ mol N}_2} \right) \left( \frac{2\text{g H}_2}{1\text{ mol H}_2} \right) = 10.7\text{g H}_2$$

3. How many grams of ammonia are produced in the reaction in Problem 2?

$$50\text{g N}_2 \quad 50\text{g} + 10.7 = \boxed{60.7\text{g NH}_3}$$



How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?

$$5.0\text{g AgNO}_3 \left( \frac{1\text{ mol}}{170\text{g}} \right) \left( \frac{1\text{ mol AgCl}}{1\text{ mol AgNO}_3} \right) \left( \frac{143\text{g}}{1\text{ mol AgCl}} \right) = 4.21\text{g AgCl}$$

5. How much barium chloride is necessary to react with the silver nitrate in Problem 4?

$$5.0\text{g AgNO}_3 \left( \frac{1\text{ mol}}{170\text{g AgNO}_3} \right) \left( \frac{1\text{ mol BaCl}_2}{2\text{ mol AgNO}_3} \right) \left( \frac{207\text{g}}{1\text{ mol BaCl}_2} \right) = \boxed{5.0\text{g BaCl}_2}$$

9-7

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