

## Unit 6 – Worksheet 3

### Ionic Compounds

#### Properties

#### Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- a.  $\text{Na}_2\text{O}$
- b.  $\text{K}_2\text{S}$
- c.  $\text{MgCl}_2$
- d.  $\text{CaBr}_2$
- e.  $\text{BaI}_2$
- f.  $\text{Al}_2\text{S}_3$
- g.  $\text{CsBr}$
- h.  $\text{AgF}$

2. Give the name of the following simple binary ionic compounds.

- a.  $\text{Na}_3\text{N}$
- b.  $\text{K}_2\text{O}$
- c.  $\text{AgBr}$
- d.  $\text{MgI}_2$
- e.  $\text{SrO}$

3. Write the formula for the following binary ionic compounds.

- a. lithium bromide
- b. sodium iodide
- c. silver sulfide
- d. cesium oxide
- d. beryllium iodide
- f. barium hydride
- g. aluminum fluoride
- h. potassium oxide

4. Write the formula for these ionic substances.

- a. silver oxide
- b. aluminum sulfide
- c. sodium nitride
- d. barium chloride
- e. strontium hydride

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**Patrick Vollmar**



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## **Table of Contents Modeling Chemistry U6 Ws 2 V2 0 Answers**

1. Understanding the eBook Modeling Chemistry U6 Ws 2 V2 0 Answers
  - The Rise of Digital Reading Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U6 Ws 2 V2 0 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Personalized Recommendations
  - Modeling Chemistry U6 Ws 2 V2 0 Answers User Reviews and Ratings
  - Modeling Chemistry U6 Ws 2 V2 0 Answers and Bestseller Lists
5. Accessing Modeling Chemistry U6 Ws 2 V2 0 Answers Free and Paid eBooks
  - Modeling Chemistry U6 Ws 2 V2 0 Answers Public Domain eBooks
  - Modeling Chemistry U6 Ws 2 V2 0 Answers eBook Subscription Services

- Modeling Chemistry U6 Ws 2 V2 0 Answers Budget-Friendly Options
- 6. Navigating Modeling Chemistry U6 Ws 2 V2 0 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry U6 Ws 2 V2 0 Answers Compatibility with Devices
  - Modeling Chemistry U6 Ws 2 V2 0 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Highlighting and Note-Taking Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Interactive Elements Modeling Chemistry U6 Ws 2 V2 0 Answers
- 8. Staying Engaged with Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry U6 Ws 2 V2 0 Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry U6 Ws 2 V2 0 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Setting Reading Goals Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Fact-Checking eBook Content of Modeling Chemistry U6 Ws 2 V2 0 Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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