

General Chemistry - Unit 1 (Part 1) Review WS - **ANSWERS**

Directions: Complete the following questions on a separate piece of paper.

LT #2: Understand the experiments and the historical theories that lead to the modern atomic theory.

- Describe Dalton's beliefs about the atom and chemical reactions with atoms.
All elements are composed of atoms; all atoms of a given element are identical, that atoms of different elements are different; chemical compounds are formed when atoms combine; atoms are not created or destroyed in chemical reactions (indivisible); solid sphere.
- Describe the atom according to Thomson's theory.
Thomson found that the atom consisted of a positive mass but there were negative pieces scattered throughout the positive area (Plum Pudding Model). He discovered the electron (negatively charged).
- Describe Rutherford's gold foil experiment. How did he change the model of the atom after the experiment?
Rutherford discovered the nucleus contained protons by shooting alpha particles (+) through gold foil. He found that most went through but some bounced back. He found the nucleus, which is very, very small and dense is found in the center of the atom, with the electrons around the outside.
- What did Chadwick change about the atomic model?
Chadwick (along with the Curies) found that the nucleus also contained the neutrons, which have approximately the same mass as the proton, but have no charge.

LT #3: Understand how subatomic particles determine the properties of an atom or ion.

- List the three subatomic particles, their location, their charge and their relative mass.
*Proton—found in nucleus, + charge, 1 amu (atomic mass unit)
Neutron—found in nucleus, no charge, 1 amu
Electron—cloud surrounding the nucleus, - charge, approx. no mass*
- What is a cation and how is it formed?
A cation is a positive ion and is formed when an atom loses one or more electrons.
 - What is an anion and how is it formed?
An anion is a negative ion and is formed when an atom gains one or more electrons.
- Complete the following charts for atoms and ions:

ATOM	ISOTOPE SYMBOL	ATOMIC #	MASS #	#PROTONS	#ELECTRONS	#NEUTRONS
Sodium - 23	²³ ₁₁ Na	11	23	11	11	12
Aluminum - 27	²⁷ ₁₃ Al	13	27	13	13	14
Chlorine - 35	³⁵ ₁₇ Cl	17	35	17	17	18
Barium - 136	¹³⁶ ₅₆ Ba	56	136	56	56	80

SYMBOL	MASS	#PROTONS	#ELECTRONS	#NEUTRONS	NET CHARGE
K ⁺¹	39	19	18	20	+1
Br ⁻¹	80	35	36	45	-1
Fe ⁺³	56	26	23	30	+3
Se ⁻²	79	34	36	45	-2

- Are all atoms of the same element identical? If not, how can they differ?
No; isotopes differ by # of neutrons. (Isotopes have the same # of protons and electrons.)
- Which of lead's four isotopes is most abundant: lead-204, lead-206, lead-207 or lead-208? Explain how you know which isotope is more abundant. *The average atomic mass for lead is 207.2 amu; therefore, the most abundant isotope is lead-207*

LT #4: Apply the mole concept (molar mass, mole equalities, conversions/dimensional analysis).

- Calculate the molar mass of each of the following atoms or compounds:
 - zinc
65.4 g/mol
 - barium phosphate, Ba₃(PO₄)₂
3(137.3) + 2(31.0) + 8(16.0) = 601.9 g/mol

Modeling Chemistry Unit 6 Review Answers

M Woodhall



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Biorefinery of Alternative Resources: Targeting Green Fuels and Platform Chemicals Sonil Nanda,Dai-Viet N. Vo,Prakash Kumar Sarangi,2020-03-31 This book summarizes recent advances in the processing of waste biomass resources to produce biofuels and biochemicals Worldwide interest in clean energy sources environmental protection and mitigating global warming is rapidly gaining momentum and spurring on the search for alternative energy sources especially for the transportation and industrial sectors This book reviews the opportunities presented by low cost organic waste materials discussing their suitability for alternative fuel and fine chemical production physicochemical characterization conversion technologies feedstock and fuel chemistry refining technologies fuel upgrading residue management and the circular economy In addition it explores applied aspects of biomass conversion by highlighting several significant thermochemical hydrothermal and biological technologies In summary the book offers comprehensive and representative descriptions of key fuel processing technologies energy conversion and management waste valorization eco friendly waste remediation biomass supply chain lifecycle assessment techno economic analysis and the circular bioeconomy *Federal Register* ,1996-07-24

Ultrasound Technologies for Food and Bioprocessing Hao Feng,Gustavo Barbosa-Canovas,Jochen Weiss,2010-11-17 Traditional food and bioprocessing technologies are facing challenges due to high expectation from the consumers and producers for better quality and safety higher process efficiency and products with novel properties or functionalities For this reason in the last few years new forms of physical energies have been explored to propose alternatives to traditional processing technologies Acoustic energy has the potential to replace or partially substitute conventional processes and at the same time offer unique opportunities in the characterization of foods and biomaterials This book is a resource for experts and newcomers in the field of power ultrasound gives insights into the physical principles of this technology details the latest advancements and links them to current and potential applications in the food and bioprocessing related industries

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