#### Chemistry Unit 7 Review

#### 701 - Inquiry Lab

- Define...
  - a. Direct relationship: Two variables increase/decrease together
  - b. Indirect relationship: One variable increases, the other decreases
- 2. Identify whether the listed variables have a direct or indirect relationship to each other.
  - a. Moles and volume: Direct

- c. Pressure and volume: Indirect
- Temperature and pressure; Direct
- d. Moles and pressure: Direct
- 3. Using the Ideal Gas Law equation, list 3 ways you could increase the volume of a balloon.

Increase the moles, increase the temperature, decrease the pressure (moreso, the pressure drops as a result of greater volume)

4. Using the Ideal Gas Law equation, list 3 ways you could decrease the pressure within a syringe.

Decrease the moles, decrease the temperature, increase the volume

5. If the number of moles in a container is doubled, by how much should the pressure increase (assuming everything else is held constant)?

The pressure should also double

#### 702 - Gas Laws Equation

1. What equation should you use when you are given just one set of variables?

PV = nRT

 You have a 4.5 L canister of O<sub>2</sub> that you store in a room that is 20 °C. The pressure of this tank is 13000 kPa. How many moles of O<sub>2</sub> are in this canister?

# **Modeling Chemistry U7 Test Key**

**Roman Wölfel** 

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## **Table of Contents Modeling Chemistry U7 Test Key**

- 1. Understanding the eBook Modeling Chemistry U7 Test Key
  - The Rise of Digital Reading Modeling Chemistry U7 Test Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Modeling Chemistry U7 Test Key
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U7 Test Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modeling Chemistry U7 Test Key
  - Personalized Recommendations
  - Modeling Chemistry U7 Test Key User Reviews and Ratings
  - Modeling Chemistry U7 Test Key and Bestseller Lists

- 5. Accessing Modeling Chemistry U7 Test Key Free and Paid eBooks
  - Modeling Chemistry U7 Test Key Public Domain eBooks
  - Modeling Chemistry U7 Test Key eBook Subscription Services
  - Modeling Chemistry U7 Test Key Budget-Friendly Options
- 6. Navigating Modeling Chemistry U7 Test Key eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Modeling Chemistry U7 Test Key Compatibility with Devices
  - Modeling Chemistry U7 Test Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U7 Test Key
  - Highlighting and Note-Taking Modeling Chemistry U7 Test Key
  - Interactive Elements Modeling Chemistry U7 Test Key
- 8. Staying Engaged with Modeling Chemistry U7 Test Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - $\circ$  Following Authors and Publishers Modeling Chemistry U7 Test Key
- 9. Balancing eBooks and Physical Books Modeling Chemistry U7 Test Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry U7 Test Key
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U7 Test Key
  - Setting Reading Goals Modeling Chemistry U7 Test Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U7 Test Key
  - Fact-Checking eBook Content of Modeling Chemistry U7 Test Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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