

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			1.2 moles KClO_3 needed
Before:	1.2	0	0	
Change:	-1.2	+1.2	+1.8	
After:	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				0.200 moles HCl needed
Before:	0.100	?	0	0	
Change:	-0.100	-0.200	+0.100	+0.100	
After:	0	?	0.100	0.100	

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
Before:	2.20	xs	0	0
Change:	-2.20	-7.70	+4.40	+6.60
After:	0	xs	4.40	6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
Before:	0.0040	XS	0	0
Change:	-0.0040	-0.0080	+0.0040	+0.0080
After:	0	XS	0.0040	0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

Modeling Chemistry Unit 13 Packet

Ying-Ying Zheng



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Table of Contents Modeling Chemistry Unit 13 Packet

1. Understanding the eBook Modeling Chemistry Unit 13 Packet
 - The Rise of Digital Reading Modeling Chemistry Unit 13 Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 13 Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 13 Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 13 Packet
 - Personalized Recommendations
 - Modeling Chemistry Unit 13 Packet User Reviews and Ratings
 - Modeling Chemistry Unit 13 Packet and Bestseller Lists

5. Accessing Modeling Chemistry Unit 13 Packet Free and Paid eBooks
 - Modeling Chemistry Unit 13 Packet Public Domain eBooks
 - Modeling Chemistry Unit 13 Packet eBook Subscription Services
 - Modeling Chemistry Unit 13 Packet Budget-Friendly Options
6. Navigating Modeling Chemistry Unit 13 Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 13 Packet Compatibility with Devices
 - Modeling Chemistry Unit 13 Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 13 Packet
 - Highlighting and Note-Taking Modeling Chemistry Unit 13 Packet
 - Interactive Elements Modeling Chemistry Unit 13 Packet
8. Staying Engaged with Modeling Chemistry Unit 13 Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 13 Packet
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 13 Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 13 Packet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 13 Packet
 - Setting Reading Goals Modeling Chemistry Unit 13 Packet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 13 Packet
 - Fact-Checking eBook Content of Modeling Chemistry Unit 13 Packet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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