

**Unit 7 Chemical Reactions**  
**PRACTICE TEST/ REVIEW**

- \_\_\_\_\_ 1. A chemical reaction is a process in which
- a. products change into reactants
  - b. the law of conservation of mass applies
  - c. substances can change state
  - d. all of these
- \_\_\_\_\_ 2. During a chemical reaction,
- a. new elements are produced
  - b. atoms are rearranged
  - c. atoms are destroyed
  - d. elements are destroyed
- \_\_\_\_\_ 3. An equation is balanced by
- a. changing subscripts
  - b. adding coefficients
  - c. erasing elements as necessary
  - d. adding elements as necessary
- \_\_\_\_\_ 4. What are the reactants in the following chemical equation:
- $$\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$$
- a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 5. What are the products in the above equation?
- a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 6. The breaking down of a substance into two or simpler substances is \_\_\_\_\_
- a. Displacement
  - b. Synthesis
  - c. Decomposition
  - d. Double displacement
- \_\_\_\_\_ 7. Each substance to the right of the arrow in a chemical equation is a(n)
- a. Reactant
  - b. Catalyst
  - c. Product
  - d. Inhibitor
- \_\_\_\_\_ 8. According to the law of conservation of mass, if two atoms of hydrogen are used as a reactant, how many atoms of hydrogen must be part of the product?
- a. 4
  - b. 2
  - c. 1
  - d. 0
- \_\_\_\_\_ 9. What type of reaction is shown in the following chemical equation,  $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$ ?
- a. Displacement
  - b. Synthesis
  - c. Decomposition
  - d. Double displacement

# Modeling Chemistry Unit 7 Chemical Reactions Key

**L Cohen**



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