

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

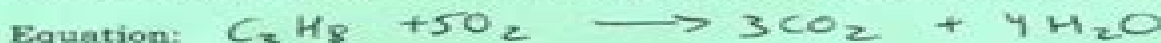


Before 8 moles $\xrightarrow{\text{excess}}$ 0 mol 0 mol

Change -8 mol -12 mol $+8 \text{ mol}$ $+8 \text{ mol}$

After $—$ excess mol O_2 8 mol SO_2 $8 \text{ mol H}_2\text{O}$

- Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?



Before $?$ mol excess 0 mol 0 mol

Change $-4 \text{ mol C}_3\text{H}_8$ -20 mol O_2 $+12 \text{ mol CO}_2$ $+16 \text{ mol H}_2\text{O}$

After $0 \text{ mol C}_3\text{H}_8$ excess mol O_2 12 mol CO_2 $16 \text{ mol H}_2\text{O}$

- Ammonia, NH_3 , for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?



Before excess mol 0.15 mol 0 mol

Change excess H_2 -0.15 mol $+0.3 \text{ mol}$

After excess H_2 0 mol N_2 0.3 mol NH_3

$$12 \text{ mol CO}_2 \times \frac{5 \text{ mol O}_2}{3 \text{ mol CO}_2}$$

$$0.15 \text{ mol N}_2 \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} = 0.3 \text{ mol NH}_3$$

Modeling Chemistry Unit 8 1

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