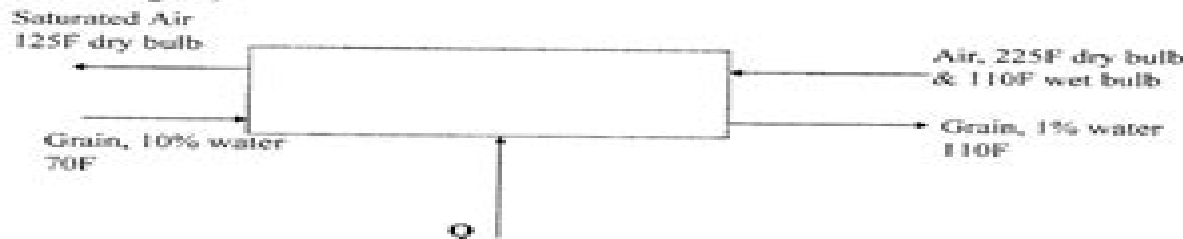


- Air at 45°C and 10% relative humidity is to be humidified adiabatically to 60% relative humidity. Use the psychrometric chart to obtain the following numbers:
  - (5 points) The dry bulb and adiabatic saturation temperature of the final air.
  - (5 points) The dew point of the final air
  - (5 points) The absolute humidity and humid volume of final air.
  - (5 points) The rate at which water must be added to humidify 15m<sup>3</sup>/min of the entering air.
- (20 points) Two aqueous sulfuric acid solutions containing 20.0 wt% H<sub>2</sub>SO<sub>4</sub> (specific gravity=1.139) and 60.0 wt % H<sub>2</sub>SO<sub>4</sub> (specific gravity=1.498) are mixed to form a 4.00 molar solution (specific gravity=1.213). Assuming 100 kg of 20% solution as the basis calculate:
  - (10 points) the mass fraction of sulfuric acid in the product solution.
  - (10 points) the feed ratio (liters of 20% solution/liter of 60% solution).
- (30 points) A rotary dryer operating at atmospheric pressure dries 10 tons/day of wet grain at 70°F, from a moisture content of 10% to 1% moisture. The air flow is countercurrent to the flow of grain, enters at 225°F dry bulb and 110°F wet-bulb temperature, and leaves at 125°F dry-bulb (see figure).



Determine: (a) the humidity of entering and leaving air if the latter is saturated, (b) the water removal in pounds per hour, (c) the daily product output in pounds per day, and (d) the heat input to the dryer. Assume that there is no heat loss from the dryer, that the grain is discharged at 110°F, and that its specific heat is 0.18.

- (30 points) A fuel gas containing 85 mole % methane and the balance ethane is burned completely with pure oxygen at 25 °C at 1atm. Suppose the reactor is continuous. Take a basis of calculation of 1 mol/s of the fuel gas, combusted with stoichiometric amount of oxygen.
  - (10 points) Please calculate the rate at which heat must be transferred from the reactor if the products are brought back down to 25 °C.
  - (15 points) Please calculate the final temperature of the combustion gas if the process is under adiabatic condition at constant pressure 1atm.
  - (5 points) Please calculate the final volume of the combustion gas if the process is under adiabatic condition at constant pressure 1atm. If the volume of the reactor can only be constant and the pressure can change, what do you expect to happen to the reactor?

# Memorandum Of Final Exam 2014

**Michael Brown**



**Memorandum Of Final Exam 2014:**

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