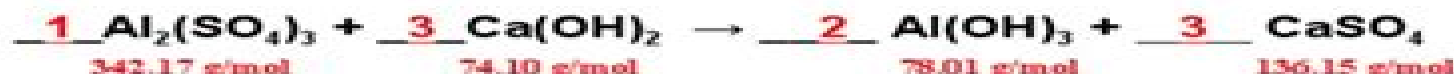


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MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.



STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: WS.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

$$\% \text{ Al} = \frac{\text{total mass Al}}{\text{molar mass Al}_2(\text{SO}_4)_3} \times 100 \% = \frac{53.96 \text{ g}}{342.17 \text{ g}} \times 100 \% = 15.77 \%$$

STEP 4: Complete the following problems using **a factor chart with units & chemicals**. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

1) How many **grams of calcium sulfate** are produced from reacting **2.59 moles of aluminum sulfate**?

$$2.59 \text{ mol Al}_2(\text{SO}_4)_3 \times \frac{3 \text{ mol CaSO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{136.15 \text{ g CaSO}_4}{1 \text{ mol CaSO}_4} = 1057.89 \text{ g CaSO}_4$$

moles A $\longrightarrow$ moles B $\longrightarrow$ grams B

Mixed Stoichiometry Practice Answers

Anthony S. Fauci



Mixed Stoichiometry Practice Answers:

Unveiling the Magic of Words: A Review of "**Mixed Stoichiometry Practice Answers**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Mixed Stoichiometry Practice Answers**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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